

Model 2 – Ternary Acids (Oxyacids)

| Acid | Name of Acid in Aqueous Solution | Cation (+) | Polyatomic Anion (-) | Polyatomic Anion Name |
|-------------------------|----------------------------------|-------------------------|----------------------|-----------------------|
| HClO_3 | Chloric acid | H_3O^+ | ClO_3^{-1} | Chlorate |
| H_2SO_3 | Sulfurous acid | $2\text{H}_3\text{O}^+$ | SO_3^{-2} | Sulfite |
| H_2SO_4 | Sulfuric acid | $2\text{H}_3\text{O}^+$ | SO_4^{-2} | Sulfate |
| H_3PO_3 | Phosphorous acid | $3\text{H}_3\text{O}^+$ | PO_3^{-3} | Phosphite |
| H_3PO_4 | Phosphoric acid | $3\text{H}_3\text{O}^+$ | PO_4^{-3} | Phosphate |
| HNO_3 | Nitric acid | H_3O^+ | NO_3^{-1} | Nitrate |
| HNO_2 | Nitrous acid | H_3O^+ | NO_2^{-1} | Nitrite |
| H_2CO_3 | Carbonic acid | $2\text{H}_3\text{O}^+$ | CO_3^{-2} | Carbonate |

6. Look at the formulas of the ternary acids in Model 2.

a. How are ternary acids different from binary acids in their structure?

Ternary acids contain three elements and they also contain the element oxygen.

b. What number do you think the prefix "ter-" refers to?

"ter" refers to three

7. When ternary acids are mixed with water, ions will form. Fill in the table above with the formulas and names of the anions.

8. Examine the pairs of ternary acids in Model 2 that contain sulfur, phosphorus, and nitrogen. Each pair has one acid that ends in "-ic" and another that ends in "-ous." These endings are related to the name of the polyatomic anion found in the acid ("-ate" or "-ite"). Complete the statements below with the correct acid name ending.

Polyatomic anion ending is "-ate" → acid name ending is "ic"

Polyatomic anion ending is "-ite" → acid name ending is "ous"

9. If the prefix "hydro-" were used to name a ternary acid, what problem would this create when naming HClO_3 ?

ClO_3^{-1} has the name of Chlorate. "ate" must be replaced with "ic". If "hydro" is used as a prefix, it would be hydrochloric acid. This has the formula HCl , not HClO_3 .

10. Write a rule for naming ternary acids.

- The prefix "hydro" is not used in the name when an acid contains more than two elements and one of the elements is oxygen.

- "ate" is replaced with "ic" and "ite" is replaced with "ous".

11. Predict the formula for chlorous acid.



12. Circle the acid(s) below that would be named beginning with the prefix "hydro-."



Polyatomic Pogil Answers

Ashok Kumar Jain, Queen's University



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Landolt-Börnstein Numerical Data and Functional Relationships in Science and Technology J. H. Callomon, K. H. Hellwege, Hans Landolt, Richard Börnstein, 1976 *Structure of Free Polyatomic Molecules* K. Kuchitsu, Koza Kuchitsu, 1998-09-10 This volume *Structure of Free Polyatomic Molecules Basic Data* contains frequently used data from the corresponding larger Landolt B rnstein handbooks in a low price book for the individual scientists working in the laboratory Directories link to the more complete volumes in the library The book contains important information about a large number of semiconductors **Structure Data of Free Polyatomic Molecules** K. Kuchitsu, 1995-11-27 Since the publication of Volumes II 7 in 1976 and its supplements II 15 in 1987 and II 21 in 1992 the information on the structure of free molecules in the ground state and in excited electronic states has increased considerably Therefore this volume II 23 contains data from 148 inorganic and 498 organic polyatomic free molecules including free radicals and molecular ions published between 1990 and 1993 inclusively and a small number of structures published 1994 All experimental methods for the determination of structural data of free molecules have been considered all data obtained by these methods have been critically evaluated and compiled The structural data for more than 3400 polyatomic free molecules can be completely surveyed and easily retrieved by means of this volume **Polyatomic Molecules** Robert S. Mulliken, Walter C. Ermler, 1981-01-28 *Polyatomic Molecules Results of Ab Initio Calculations* describes the symmetry of polyatomic molecules in ground states This book contains 12 chapters that also cover the excited and ionized states of these molecules The opening chapter describes the nature of the various ab initio computational methods The subsequent four chapters deal with the three atom systems differing with respect to the number of hydrogen atoms in the molecules These chapters also discuss the reaction surfaces of these systems These topics are followed by discussions on the molecules whose ground states belong to relatively high little or no symmetry groups The concluding chapters explore the inorganic and relatively large organic molecules These chapters also examine the ab initio calculations of molecular compounds and complexes as well as hydrogen bonding and ion hydration This text will be of great value to organic and inorganic chemists and physicists Polyatomic Ion-surface Interactions Luke Hanley, 1998 **The Particulate Nature of Polyatomic Ions** ,2007 *Theoretical Investigations of the Shapes of Simple Polyatomic Molecules* Yuuzi Takahata, 1970 The Entropy of Polyatomic Molecules and the Symmetry Number Joseph Edward Mayer, Stephen Brunauer, Maria Goeppert Mayer, 1933 The Geometric Phase in Polyatomic Molecules Mark Steven Holt, 2012 **Polyatomic Molecules for Probing a Possible Variation of the Proton-to-electron Mass Ratio** Paul Jansen, 2013 On the Born-Oppenheimer Expansion for Polyatomic Molecules M. Klein, A. Martinez, Université Paris-Nord. Département de mathématique et informatique, R. Seiler, 1990 Approximate Molecular-orbital Calculations on Polyatomic Molecules Bruce Gilson, 1969 **Nonequilibrium phenomena in polyatomic gases** Frederick R. MacCourt, **Nonequilibrium Phenomena in Polyatomic Gases** ,1990 **Polyatomic Molecules Without Internal Rotation**

,1968 **The Scattering of Slow Electrons and Positrons by Polyatomic Molecules** Ashok Kumar Jain, Queen's University, 1983 Polyatomic Molecules with Internal Rotation Paul F. Wacker, 1969 *The Production and Spectroscopy of Small Polyatomic Molecular Ions Isolated in Solid Neon*, 1990 The Rotational Ramon Effect in Polyatomic Gases Charlton M. Lewis, California Institute of Technology. Division of Physics, Mathematics and Astronomy, 1933 **Some transport phenomena in dilute and rarefied polyatomic gases** Jacob Nicolaas Breunese, 1982

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