

Answer following



- 1. Substrate and temperature:  $20^{\circ}\text{C}$   $30^{\circ}\text{C}$   $40^{\circ}\text{C}$
- 2. A graph of initial velocity:  $100\text{ mg/ml}$   $200\text{ mg/ml}$   $300\text{ mg/ml}$
- 3. A graph of initial velocity:  $100\text{ mg/ml}$   $200\text{ mg/ml}$   $300\text{ mg/ml}$

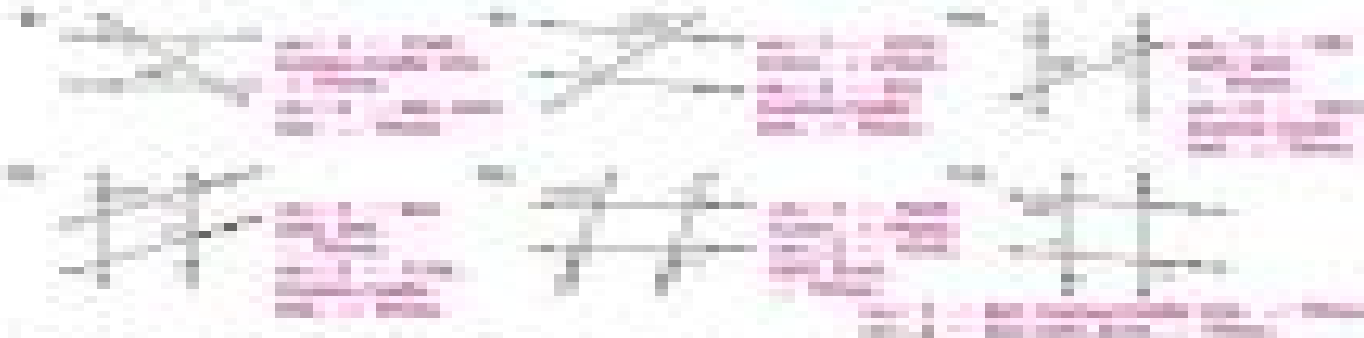
Calculate all values of  $K_m$  and  $V_{max}$  for the following reaction. (Note:  $K_m$  is the substrate concentration at which the reaction rate is half of  $V_{max}$ .)

- 1.  $K_m = 100\text{ mg/ml}$   $V_{max} = 100\text{ mg/ml}$
- 2.  $K_m = 100\text{ mg/ml}$   $V_{max} = 200\text{ mg/ml}$
- 3.  $K_m = 100\text{ mg/ml}$   $V_{max} = 300\text{ mg/ml}$
- 4.  $K_m = 100\text{ mg/ml}$   $V_{max} = 400\text{ mg/ml}$



Answer: 100

Graph the following data. (Note:  $K_m$  is the substrate concentration at which the reaction rate is half of  $V_{max}$ .)



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**Michael Hoelscher**



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