

pH AND pOH

Name _____

The pH of a solution indicates how acidic or basic that solution is.

- pH range of 0 - 7 acidic
- 7 neutral
- 7-14 basic

Since $[H^+][OH^-] = 10^{-14}$ at $25^\circ C$, if $[H^+]$ is known, the $[OH^-]$ can be calculated and vice versa.

$$pH = -\log [H^+]$$

$$\text{So if } [H^+] = 10^{-6} M, pH = 6.$$

$$pOH = -\log [OH^-]$$

$$\text{So if } [OH^-] = 10^{-8} M, pOH = 8.$$

$$\text{Together, } pH + pOH = 14.$$

Complete the following chart.

	$[H^+]$	pH	$[OH^-]$	pOH	Acidic or Basic
1.	$10^{-5} M$	5	$10^{-9} M$	9	Acidic
2.		7			
3.			$10^{-4} M$		
4.	$10^{-2} M$				
5.				11	
6.		12			
7.			$10^{-5} M$		
8.	$10^{-11} M$				
9.				13	
10.		6			

Ph Review Problems Worksheet Answers

Deanna Davis Ph.D.



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