
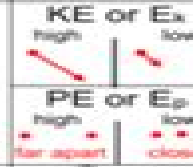





Study Buddy: Energy and Phase Changes

Terms	Description	Picture	Symbol/Unit
Heat	Amount of thermal energy transferred from one system to another. <small>http://openstax.org/r/media_img/video/FILE:Heat1_1st.png http://www.ck12.org/ck12-quartz/277502/277502-cv-0001-mid.png</small>		Q/ Joules (J) or calorie (cal)
Total Internal Energy	Describes the total amount of energy in a system. Equals the sum of kinetic and potential energies. Kinetic Energy: due to motion of particles Potential Energy: due to relative position of particles		IE/ Joules (J) or calorie (cal)
Endothermic Process	System absorbs energy from surroundings Phase changes: melting, vaporizing, subliming <small>http://openstax.org/r/cooling-curve-plot-creating-page-2015</small>		+Q for system
Exothermic Process	System releases heat to surroundings Phase changes: freezing, condensing, depositing		-Q for system
Vapor Pressure	Pressure exerted by the vapor particles over a liquid or solid substance at equilibrium.		<small>http://www.ck12.org/ck12-quartz/277502/277502-cv-0001-mid.png</small>

How does thermal energy transfer cause phase changes?


As a solid or liquid is heated, the thermal energy is used to increase the kinetic and potential energy of the particles (particles moving faster and farther apart). When a liquid or gas is cooled, the thermal energy transferred to the surroundings reduced the kinetic and potential energies of the particles.



Describe solid particles' positions and motion:

- Moving slowly
- Close together

Solid water



Describe liquid particles' positions and motion:

- Moving faster
- Farther apart

Liquid water



Describe gaseous particles' positions and motion:

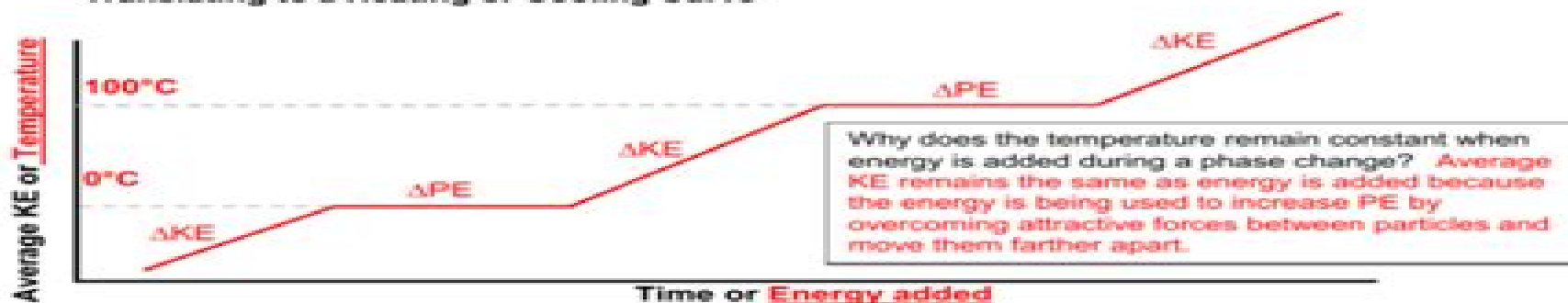
- Moving quickly
- Far apart

Water vapor

Melting = 0 °C

Boiling = 100 °C

Translating to a Heating or Cooling Curve



Label where PE and KE are changing on the heating curve above.

Phase Changes Study Guide

Rachel Harrison



Phase Changes Study Guide:

Physics for Scientists and Engineers Study Guide Todd Ruskell, 2007-06-08 The Study Guide provides students with key physical quantities and equations misconceptions to avoid questions and practice problems to gain further understanding of physics concepts and quizzes to test student knowledge of chapters *Chemistry, Student Study Guide* James E.

Brady, Fred Senese, 2008-01-28 The image on the front cover depicts a carbon nanotube emerging from a glowing plasma of hydrogen and carbon as it forms around particles of a metal catalyst Carbon nanotubes are a recently discovered allotrope of carbon Three other allotropes of carbon buckyballs graphite and diamond are illustrated at the left as is the molecule methane CH_4 from which nanotubes and buckyballs can be made The element carbon forms an amazing number of compounds with structures that follow from simple methane found in natural gas to the complex macromolecules that serve as the basis of life on our planet The study of chemistry also follows from the simple to the more complex and the strength of this text is that it enables students with varied backgrounds to proceed together to significant levels of achievement

Chemical Principles Student's Study Guide & Solutions Manual John Krenos, Joseph Potenza, Loretta Jones, Lynn Koplitz, Thomas Spence, 2004-03-19 This combination manual is designed to help students avoid common mistakes and understand the material better The solutions manual section includes detailed answers and explanations to the odd numbered exercises in the text **EPA 608 Study Guide** Mike L Turner, 2024-08-31 This comprehensive guide is designed for anyone preparing for the EPA 608 Certification exam offering an in depth exploration of key concepts and practical insights essential for passing the test Whether you are new to the HVAC field or looking to refresh your knowledge this book provides a thorough understanding of refrigerant handling safety procedures and regulatory requirements The book begins with an overview of the EPA 608 Certification outlining the importance of this credential for HVAC professionals and detailing the different types of certification including Type I Type II Type III and Universal Each chapter is meticulously crafted to cover specific aspects of the certification process ensuring you are well prepared for every section of the exam Chapter by chapter you will delve into critical topics such as EPA regulations refrigerant types and their environmental impact and safety and handling procedures The book also provides a detailed breakdown of refrigerant recovery recycling and reclamation processes equipping you with the knowledge needed to handle various refrigerants responsibly and in compliance with regulations In addition to theoretical knowledge practical insights are provided through extensive practice questions and answers tailored to each certification type These practice questions are designed to help you assess your understanding and improve your test taking strategies Detailed explanations accompany each answer offering valuable insights into common pitfalls and effective problem solving techniques Special attention is given to exam strategies including time management tips and understanding the question format to ensure you approach the exam with confidence The book also addresses common pitfalls and offers last minute preparation tips to maximize your performance on the exam day This

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The Practice of Chemistry Study Guide & Solutions Manual Pamela Mills,Amina El-Ashmawy,2003-04-14 Designed to help students understand the material better and avoid common mistakes Also includes solutions and explanations to odd numbered exercises

Chemical Principles Study Guide/Solutions Manual John Krenos,Joseph Potenza,Carl Hoeger,2007-01-18 Written for general chemistry courses Chemical Principles helps students develop chemical insight by showing the connection between chemical principles and their applications

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