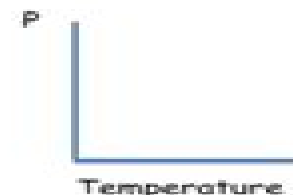
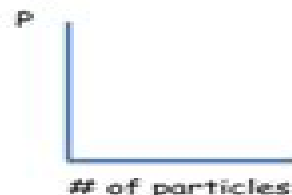
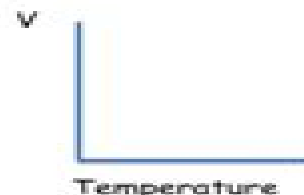


## Bouncing off the Walls PhET Gas Law Simulation 2013

**GO TO:** <http://phet.colorado.edu/en/simulation/gas-properties> OR google "Phet Gas properties".

**GOAL:** to investigate the effect of: a) volume on pressure; b) temperature on volume;  
c) number of particles on pressure; d) temperature on pressure.

**PRELAB:** Sketch your prediction of each relationship.



### COLLECTING DATA

#### PART 1: Effect of Volume on Pressure

1. Hit the reset button.
2. In the upper left corner under "Constant Parameters", select "Temperature". This will keep the temperature relatively constant. Set the number of particles to a constant value.
3. To measure the volume, hit the button "Measurement Tools" and select the "Ruler". The ruler can be moved so you can measure the size of the box to get a measurement of volume.
4. Take 5 separate measurements over a large range of numbers without exploding the container.

**Table 2:** Volume and Pressure of a gas with temperature constant at \_\_\_\_\_ and Particles constant at \_\_\_\_\_

Volume (nm)					
Pressure (atm)					

Questions:

1. What is the independent variable?
2. What is the dependent variable?
3. What 2 factors are being held constant?

# Phet Exploring Gas Law Answers

**Yi-Tong Ma**

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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### **Table of Contents Phet Exploring Gas Law Answers**

1. Understanding the eBook Phet Exploring Gas Law Answers
  - The Rise of Digital Reading Phet Exploring Gas Law Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Phet Exploring Gas Law Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Phet Exploring Gas Law Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Phet Exploring Gas Law Answers

- Personalized Recommendations
  - Phet Exploring Gas Law Answers User Reviews and Ratings
  - Phet Exploring Gas Law Answers and Bestseller Lists
5. Accessing Phet Exploring Gas Law Answers Free and Paid eBooks
    - Phet Exploring Gas Law Answers Public Domain eBooks
    - Phet Exploring Gas Law Answers eBook Subscription Services
    - Phet Exploring Gas Law Answers Budget-Friendly Options
  6. Navigating Phet Exploring Gas Law Answers eBook Formats
    - ePub, PDF, MOBI, and More
    - Phet Exploring Gas Law Answers Compatibility with Devices
    - Phet Exploring Gas Law Answers Enhanced eBook Features
  7. Enhancing Your Reading Experience
    - Adjustable Fonts and Text Sizes of Phet Exploring Gas Law Answers
    - Highlighting and Note-Taking Phet Exploring Gas Law Answers
    - Interactive Elements Phet Exploring Gas Law Answers
  8. Staying Engaged with Phet Exploring Gas Law Answers
    - Joining Online Reading Communities
    - Participating in Virtual Book Clubs
    - Following Authors and Publishers Phet Exploring Gas Law Answers
  9. Balancing eBooks and Physical Books Phet Exploring Gas Law Answers
    - Benefits of a Digital Library
    - Creating a Diverse Reading Collection Phet Exploring Gas Law Answers
  10. Overcoming Reading Challenges
    - Dealing with Digital Eye Strain
    - Minimizing Distractions
    - Managing Screen Time
  11. Cultivating a Reading Routine Phet Exploring Gas Law Answers
    - Setting Reading Goals Phet Exploring Gas Law Answers
    - Carving Out Dedicated Reading Time
  12. Sourcing Reliable Information of Phet Exploring Gas Law Answers

- Fact-Checking eBook Content of Phet Exploring Gas Law Answers
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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