

Name : _____ Date : _____



pH and pOH Worksheet



1. A solution has a hydrogen ion concentration of 2.8×10^{-6} M. What is the pOH of its solution ?

2. A solution has a $[\text{OH}^-]$ of 5.8×10^{-7} . What is the pH of this solution ?

3. A 450 mL beaker is 0.00045 M HCl. What is the pH of this solution ?

4. A 320 mL solution contains 2.30 mg of NaOH. What is the pH of this solution ?

5. Complete the following table.

pH	$[\text{H}^+]$	$[\text{OH}^-]$	pOH
3.5			
	5.8×10^{-7}		
		4.2×10^{-2}	
			8.2
	4.2×10^{-5}		
			2.4
10.1			
		7.2×10^{-3}	

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developing field of environmental chemistry Intense one semester approach to learning Application based approach to learning theoretical concepts In depth analysis of field based and in situ analytical techniques Introduction to environmental modeling

Chemistry and Physics for Nurse Anesthesia, Second Edition David Shubert, David C Shubert, John Leyba, 2013-03-15 Praise for the first edition A welcome addition to the reference materials necessary for the study of nurse anesthesia The textbook is divided into logical easy to use sections that cover all areas necessary for the practice of nurse anesthesia This is a text that is easy to read and able to be incorporated into any nurse anesthesia chemistry and physics course I would recommend this textbook to any program director Anthony Chipas PhD CRNA Division Director Anesthesia for Nurses Program Medical University of South Carolina Nurse anesthesia students will welcome the second edition of this text designed for the combined course in chemistry and physics that is required for this program It is written in a clear conversational style to counteract the trepidation that often accompanies the study of chemistry and physics and includes only those core scientific concepts that relate to clinical anesthesia application Numerous illustrations demonstrate how the scientific concepts relate directly to their clinical application in anesthesia and plentiful case studies exemplify and reinforce basic concepts Review question at the end of each chapter facilitate self assessment This second edition offers numerous features that will further assist students with understanding and mastery of the material These new features are the direct result of knowledge gained from on line and traditional classroom teaching experiences They include chapter summaries additional questions and answers at the end of each chapter specific to nurse anesthesia end of chapter summaries and lists of formulas and constants discussed in the book Fifteen videos vividly demonstrate the key principles of the chemistry and physics of nurse anesthesia Corresponding to various sections of the book they supplement and illustrate text content Also available are revised PowerPoint slides for faculty use The first edition of this popular text is currently being used by eight nurse anesthesia programs throughout the United States and many additional programs plan to adopt the second edition New to the Second Edition Emphasizes content in chemistry and physics that relates specifically to anesthesia with a strong focus on gases Includes case studies to illustrate and reinforce knowledge Provides additional end of chapter problems focused on anesthesia Relates core scientific concepts to clinical anesthesia application Offers fifteen videos demonstrating key principles of the physics and chemistry of nurse anesthesia

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CliffsNotes AP Chemistry Bobrow Test Preparation Services,2009-02-09 The book itself contains chapter length subject reviews on every subject tested on the AP Chemistry exam as well as both sample multiple choice and free response questions at each chapter s end Two full length practice tests with detailed answer explanations are included in the book

Fundamentals of Chemistry Ralph A. Burns,1995

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