

KEY

Unit 9 Outline- Thermodynamics, Reaction Rates, & Chemical Equilibrium Chapters 17 & 18

Essential Skills/ State Standards:

- 9-1. Describe how temperature and heat flow in terms of the motion of molecules (or atoms). (CSS: 7a)
 9-2. Know that chemical processes & phase changes can either release energy (exothermic) or absorb energy (endothermic). (CSS: 7b, 7c)
 9-3. Solve problems involving heat flow and temperature changes, using known values of specific heat ($q = m \cdot c \cdot \Delta T$) & latent heat of phase change ($q = m \cdot \Delta H$). (CSS: 7d)
 9-4. Know the rate of reaction is the decrease in concentration of reactants or the increase in concentration of products with time & depends on such factors as: concentration, temperature, and pressure. (CSS: 8a, 8b)
 9-5. Know the definition and role of activation energy & catalyst in a chemical reaction. (CSS: 8d*, 8c)
 9-6. Know how to use Le Chatelier's principle to predict the effect of changes in concentration, temperature, and pressure. (CSS: 9a)
 9-7. Know that equilibrium is established when the forward and reverse reaction rates are equal. (CSS: 9b)

Vocabulary

- | | | | |
|----------------|-------------------|--|---------------------------|
| • Condensation | • Deposition | • Calorimeter | • Heat of reaction |
| • Freezing | • Heat | • Joule | • Le Chatlier's principle |
| • Vaporization | • Temperature | • Specific heat | • Equilibrium |
| • Melting | • Thermochemistry | • $q = m \cdot c \cdot \Delta T$ | • Activation Energy |
| • Boiling | • Exothermic | • $q = m \cdot \Delta H_{\text{vaporization}}$ | • Catalyst |
| • Sublimation | • Endothermic | • $q = m \cdot \Delta H_{\text{fusion}}$ | |

Study Guide (Use notes, worksheets, and labs in addition to this study guide)

1. According to the collision theory, what 3 things must happen in order for two molecules to react?
 Must collide with each other, Must collide with enough energy, & in the correct orientation

2. What is the importance of the activation energy in a chemical reaction?
 Amount of energy needed for a reaction to occur.

3. Using the collision theory, explain if the following would increase/ decrease the rate of reaction:
 $2 \text{HCl(aq)} + \text{Mg(s)} \rightarrow \text{MgCl}_2\text{(aq)} + \text{H}_2\text{(g)}$

- a. Increase temp: Increase, molecules collide more
 b. Increase HCl: Increase, more collisions because more reactants to collide
 c. Increase pressure: Hard to say, wouldn't affect the reactants directly because none of them are gases.

One of the products is a gas, so it might cause the gas particles to collide more with the reactants and could slow it down. Tough call.. (this one won't be on the test).

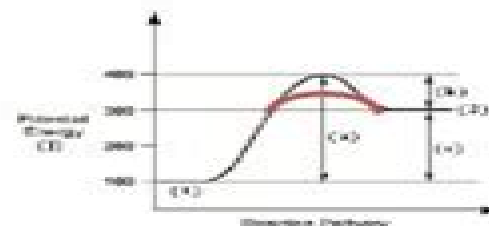
- d. Remove MgCl_2 : Increase, less products means less chance for them to interfere with reactant collisions
 e. Add H_2 : Decrease, more products mean they interfere more with reactants and the reactants would collide less.

4. Identify the following terms with their correct letter from the graphs:

- * Reactants- 1
 * Products- 2
 * Activation energy (number/ letter & amount)-
 A (300 Joules)
 * Is this reaction exothermic or endothermic?

Endothermic

5. Explain how a catalyst affects the activation energy in a chemical reaction. (Also, show on the above graph what a catalyst might do). It lowers the activation energy (see on graph).
 Speeds it up by lowering the activation energy.



Reaction Rates Equilibrium Test Answers

R Bogdan



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