

## Calculating Average Rates of Reaction

*You are required to show work for all calculations.*

1. During the combustion of methane,  $\text{CH}_4$ , shown by the reaction



The concentration of methane was measured at various time intervals and the following results were obtained:

Time (s)	$[\text{CH}_4]$ ( $\text{mol} \cdot \text{l}^{-1}$ )
10	2.40
20	1.20
30	0.80
40	0.60

Calculate the average rate of loss of methane during the 10 to 40 second time period.

**Solution:**

$$\begin{aligned}
 \text{rate} &= \frac{\Delta[\text{CH}_4]}{\Delta t} \\
 &= \frac{[0.60 - 2.4]\text{mol}}{(40 - 10)\text{s} \times \text{L}} \\
 &= \frac{-1.8\text{mol}}{30\text{s} \times \text{L}} \\
 &= -0.06 \frac{\text{mol}}{\text{L} \times \text{s}}
 \end{aligned}$$

Note: the answer is negative because we are calculating the **rate of loss of methane**.

## Calculating Reaction Rates using Stoichiometry

**Example:**

In the following decomposition reaction,



# Reaction Rates Practice Problems Answers

**Siddappa N.Byrareddy**



## Reaction Rates Practice Problems Answers:

(Free Sample) *GO TO Objective NEET Chemistry Guide with DPP & CPP Sheets 9th Edition* Disha Experts, 2021-10-07

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*Survival Guide to General Chemistry* Patrick E. McMahon, Rosemary McMahon, Bohdan Khomtchouk, 2019-02-13. This work evolved over thirty combined years of teaching general chemistry to a variety of student demographics. The focus is not to recap or review the theoretical concepts well described in the available texts. Instead, the topics and descriptions in this book make available specific detailed step-by-step methods and procedures for solving the major types of problems in general chemistry. Explanations, instructional process sequences, solved examples, and completely solved practice problems are greatly expanded, containing significantly more detail than can usually be devoted to in a comprehensive text. Many chapters also provide alternative viewpoints as an aid to understanding. Key Features: The authors have included every major topic in the first semester of general chemistry and most major topics from the second semester. Each is written in a specific and detailed step-by-step process for problem solving, whether mathematical or conceptual. Each topic has greatly expanded examples and solved practice problems containing significantly more detail than found in comprehensive texts. Includes a chapter designed to eliminate confusion concerning acid-base reactions, which often persists through working with acid-base equilibrium. Many chapters provide alternative viewpoints as an aid to understanding. This book addresses a very real need for a large number of incoming freshmen in STEM fields. **Chemistry** Rob Ritchie, 2004.

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also suited for students majoring in other fields where it is important to have an understanding of the basics of chemistry. Students need have no previous background in chemistry but should possess basic math skills. The text features numerous helpful problems and learning features.

**Survival Guide to Organic Chemistry** Patrick E. McMahon, Bohdan B. Khomtchouk, Claes Wahlestedt, 2016-12-19

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