# TYPES OF CHEMICAL REACTIONS

Complete each of the following equations as needed to make it the type of reaction indicated. Be sure to write each formula correctly.

- Double replacement: Na<sub>2</sub>CO<sub>4</sub> + PbCl 2 NaCl + PbCO<sub>4</sub>
- 2) Single replacement: Ch + 2 NaBr -> 2 NaCl + Br<sub>2</sub>

answer key

- 3) Synthesis:  $2 H_2 + 0_2 2 HO$
- 4) Double replacement: 3 Ca(OH)<sub>2</sub> + 2 FeCl 2 Fe(OH)<sub>2</sub> + 3 CaCl<sub>2</sub>
- 5) Double replacement: AgNO<sub>3</sub> + KI -> AgI + KNO<sub>3</sub>
- 6) Single replacement: Cu + 2 AgNO<sub>3</sub> -> 2 Ag + Cu(NO<sub>3</sub>)<sub>2</sub> [Copper (II) is used here]
- 7) Decomposition: Mg(ClO<sub>3</sub>)<sub>2</sub> MgCl<sub>2</sub> + 3 0<sub>2</sub>
- 8) Single replacement: Fe + Cu(NO<sub>3</sub>)<sub>2</sub> -> Fe(NO<sub>3</sub>)<sub>2</sub> + Cu [Assume Fe<sup>2\*</sup>]
- 9) Synthesis: S+02 SO2
- 10) Decomposition: 2 Hg0 4 Hg + 0

For each of the following equations, identify what kind of reaction it represents: double replacement, single replacement, decomposition, or synthesis (composition).

- 1) 2Mg + 0<sub>2</sub> 2 MgO Synthesis
- 2) Cal<sub>2</sub> + Cl<sub>2</sub> CaCl<sub>2</sub> In Single Replacement
- 3) 3KOH + AICl<sub>3</sub> -> AK(OH)<sub>3</sub> + 3KCl Double Replacement
- 4) C + O2 -> CO2 Synthesis

# **Reaction Types Answer Key**

Robin Donohoe Dennison, Kathleen Farrell

#### **Reaction Types Answer Key:**

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