

Equilibrium Practice Problems

1. Write the equilibrium expression for each of the following reactions:



$$K = \frac{[\text{NH}_3]^2}{[\text{N}_2] [\text{H}_2]^3}$$



$$K = \frac{[\text{ICl}]^2}{[\text{Cl}_2]}$$



$$K = \frac{[\text{NO}] [\text{O}_2]^{1/2}}{[\text{NO}_2]}$$

2. The dissociation of acetic acid, CH_3COOH , has an equilibrium constant at 25°C of 1.8×10^{-5} . The reaction is



If the equilibrium concentration of CH_3COOH is 0.46 moles in 0.500 L of water and that of CH_3COO^- is 8.1×10^{-3} moles in the same 0.500 L, calculate $[\text{H}^+]$ for the reaction.

$$K = \frac{[\text{CH}_3\text{COO}^-] [\text{H}^+]}{[\text{CH}_3\text{COOH}]} \quad \text{Plug in known values and solve.}$$

$$1.8 \times 10^{-5} = \frac{[8.1 \times 10^{-3} \text{ moles}/0.500 \text{ L}] [\text{H}^+]}{[0.46 \text{ moles}/0.500 \text{ L}]}$$

$$[\text{H}^+] = 1.0 \times 10^{-3} \text{ M}$$

3. Indicate the effect of a catalyst, pressure, temperature and concentration on each of the following on:

	catalyst	pressure	temperature	concentration
a. speed	presence increase	direct relationship	direct relationship	direct relationship
b. pos ⁿ	nothing	based on moles	based on enthalpy of rxn	inverse relationship

Reaction Rates And Equilibrium Practice Problems

Tracy Poulsen



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Table of Contents Reaction Rates And Equilibrium Practice Problems

1. Understanding the eBook Reaction Rates And Equilibrium Practice Problems
 - The Rise of Digital Reading Reaction Rates And Equilibrium Practice Problems
 - Advantages of eBooks Over Traditional Books
2. Identifying Reaction Rates And Equilibrium Practice Problems
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Reaction Rates And Equilibrium Practice Problems
 - User-Friendly Interface
4. Exploring eBook Recommendations from Reaction Rates And Equilibrium Practice Problems
 - Personalized Recommendations
 - Reaction Rates And Equilibrium Practice Problems User Reviews and Ratings
 - Reaction Rates And Equilibrium Practice Problems and Bestseller Lists
5. Accessing Reaction Rates And Equilibrium Practice Problems Free and Paid eBooks

- Reaction Rates And Equilibrium Practice Problems Public Domain eBooks
 - Reaction Rates And Equilibrium Practice Problems eBook Subscription Services
 - Reaction Rates And Equilibrium Practice Problems Budget-Friendly Options
6. Navigating Reaction Rates And Equilibrium Practice Problems eBook Formats
 - ePub, PDF, MOBI, and More
 - Reaction Rates And Equilibrium Practice Problems Compatibility with Devices
 - Reaction Rates And Equilibrium Practice Problems Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Reaction Rates And Equilibrium Practice Problems
 - Highlighting and Note-Taking Reaction Rates And Equilibrium Practice Problems
 - Interactive Elements Reaction Rates And Equilibrium Practice Problems
 8. Staying Engaged with Reaction Rates And Equilibrium Practice Problems
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Reaction Rates And Equilibrium Practice Problems
 9. Balancing eBooks and Physical Books Reaction Rates And Equilibrium Practice Problems
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Reaction Rates And Equilibrium Practice Problems
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Reaction Rates And Equilibrium Practice Problems
 - Setting Reading Goals Reaction Rates And Equilibrium Practice Problems
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Reaction Rates And Equilibrium Practice Problems
 - Fact-Checking eBook Content of Reaction Rates And Equilibrium Practice Problems
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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