

**Pre Public Examination**  
GCSE Mathematics (Edexcel style)  
May 2017  
Higher Tier  
**Paper 2H**

Name .....

Class .....

**TIME ALLOWED**

1 hour 30 minutes

**INSTRUCTIONS TO CANDIDATES**

- Answer **all** the questions.
- Read each question carefully. Make sure you know what you have to do before starting your answer.
- **You are permitted to use a calculator in this paper.**
- Do all rough work in this book.

**INFORMATION FOR CANDIDATES**

- The number of marks is given in brackets ( ) at the end of each question or part question on the Question Paper.
- **You are reminded of the need for clear presentation in your answers.**
- The total number of marks for this paper is **80**.

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Question	Mark	Out of
1		2
2		4
3		5
4		2
5		2
6		4
7		4
8		4
9		5
10		4
11		5
12		4
13		3
14		4
15		4
16		4
17		4
18		4
19		6
20		6
Total		80

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## **Table of Contents Pixl Jan 2014 Ppe Paper 1 Maths**

1. Understanding the eBook Pixl Jan 2014 Ppe Paper 1 Maths
  - The Rise of Digital Reading Pixl Jan 2014 Ppe Paper 1 Maths
  - Advantages of eBooks Over Traditional Books
2. Identifying Pixl Jan 2014 Ppe Paper 1 Maths
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Pixl Jan 2014 Ppe Paper 1 Maths
  - User-Friendly Interface
4. Exploring eBook Recommendations from Pixl Jan 2014 Ppe Paper 1 Maths
  - Personalized Recommendations
  - Pixl Jan 2014 Ppe Paper 1 Maths User Reviews and Ratings
  - Pixl Jan 2014 Ppe Paper 1 Maths and Bestseller Lists
5. Accessing Pixl Jan 2014 Ppe Paper 1 Maths Free and Paid eBooks

- Pixl Jan 2014 Ppe Paper 1 Maths Public Domain eBooks
- Pixl Jan 2014 Ppe Paper 1 Maths eBook Subscription Services
- Pixl Jan 2014 Ppe Paper 1 Maths Budget-Friendly Options
- 6. Navigating Pixl Jan 2014 Ppe Paper 1 Maths eBook Formats
  - ePub, PDF, MOBI, and More
  - Pixl Jan 2014 Ppe Paper 1 Maths Compatibility with Devices
  - Pixl Jan 2014 Ppe Paper 1 Maths Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Pixl Jan 2014 Ppe Paper 1 Maths
  - Highlighting and Note-Taking Pixl Jan 2014 Ppe Paper 1 Maths
  - Interactive Elements Pixl Jan 2014 Ppe Paper 1 Maths
- 8. Staying Engaged with Pixl Jan 2014 Ppe Paper 1 Maths
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Pixl Jan 2014 Ppe Paper 1 Maths
- 9. Balancing eBooks and Physical Books Pixl Jan 2014 Ppe Paper 1 Maths
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Pixl Jan 2014 Ppe Paper 1 Maths
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Pixl Jan 2014 Ppe Paper 1 Maths
  - Setting Reading Goals Pixl Jan 2014 Ppe Paper 1 Maths
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Pixl Jan 2014 Ppe Paper 1 Maths
  - Fact-Checking eBook Content of Pixl Jan 2014 Ppe Paper 1 Maths
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development

- Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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web solution 1 the two half reactions balanced as if in acidic solution  $2\text{CN}^{-} + \text{Au} + \text{CN}^{-} + 2\text{e}^{-} \rightarrow 2\text{HCN} + 2\text{OH}^{-}$   $2\text{H}_2\text{O}_2 \rightarrow 2\text{H}_2\text{O} + \text{O}_2$  make electrons equal convert to basic solution  $4\text{CN}^{-} + 2\text{Au} + 2\text{Au} + \text{CN}^{-} + 2\text{e}^{-}$  multiplied by a factor of 2  $2\text{e}^{-} + 2\text{H}_2\text{O} \rightarrow \text{O}_2 + 2\text{OH}^{-}$  3 the final answer

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web we can balance it for charge by adding two electrons to the right side of the equation so that the net charge on each side is 0 oxidation ni s ni<sup>2+</sup> a 2 e now that the oxidation half reaction is balanced it tells us that two electrons are produced for every atom of nickel oxidized but where do those electrons go

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web in the redox equation below that occurs in our stomach which of the following pairs identifies the reducing and oxidizing agents respectively  $\text{C}_6\text{H}_8\text{O}_6$   $2\text{H}_2\text{NO}_2$   $\text{C}_6\text{H}_6\text{O}_6$   $2\text{H}_2\text{O}_2$   $2\text{NO}$  choose 1 answer  $\text{C}_6\text{H}_8\text{O}_6$  and  $\text{NO}_2$  a  $\text{C}_6\text{H}_8\text{O}_6$  and  $\text{NO}_2$  no 2 and  $\text{C}_6\text{H}_8\text{O}_6$  b  $\text{NO}_2$  and  $\text{C}_6\text{H}_8\text{O}_6$  h and  $\text{NO}_2$  c h and

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the document 3 line spacing incorrect 4 letter spaces incorrect 5 upper case and or lower case incorrect 6 font and or font  
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