

$$M = \frac{\# \text{ of moles}}{V (\text{in L})}$$

Molar Concentration

Part A: Describing Molar Concentration

1. Sodium chloride was dissolved in water to produce a 1.5M solution. Explain what this concentration tells us about the NaCl solution. How might a chemist use this ratio?

$$\frac{1.5 \text{ mol of NaCl}}{1 \text{ L of water}}$$

if you need a certain amount of NaCl you can take a certain volume of the solution.

2. A 45.3 g sample of potassium nitrate is dissolved in enough water to make 225 mL of solution. Determine the molar concentration of the potassium nitrate.

$$\begin{aligned} \text{KNO}_3 : \text{m.m.} &= 101.11 \text{ g/mol} & 0.45 \text{ mol KNO}_3 \\ V &= 0.225 \text{ L} & \frac{0.45 \text{ mol KNO}_3}{0.225 \text{ L}} = 2 \text{ M} \end{aligned}$$

3. Find the molarity of a solution made from 275 g of CuSO_4 dissolved in enough water to make 4.25 L.

$$\begin{aligned} 275 \text{ g CuSO}_4 & \quad \frac{275 \text{ g}}{159.65} = 1.72 \text{ mol CuSO}_4 \\ \text{m.m.} & \quad \frac{1.72 \text{ mol}}{4.25 \text{ L}} = 0.40 \text{ M CuSO}_4 \\ 63.55 & \\ + 32.1 & \\ \hline 95.65 & \end{aligned}$$

4. An alcoholic iodine solution ("tincture" of iodine) is prepared by dissolving 5.15 g of iodine crystals in enough alcohol to make a volume of 225 mL. Calculate the molarity of iodine in the solution.

$$\begin{aligned} 5.15 \text{ g I}_2 & \quad \frac{0.02 \text{ mol I}_2}{0.225 \text{ L}} = 0.089 \text{ M I}_2 \\ 253.8 & \quad \frac{5.15 \text{ g I}_2}{253.8} = 0.02 \text{ mol I}_2 \end{aligned}$$

5. What final volume would be needed in order to prepare a 0.25 M NaCl solution from 5.2 g of NaCl (s)?

$$\begin{aligned} \text{m.m. NaCl} : & \quad \frac{5.2 \text{ g}}{58.44 \text{ g/mol}} = 0.089 \text{ mol NaCl} & M = \frac{\# \text{ of mol}}{V (\text{in L})} \\ 22.99 & & V (\text{in L}) = \frac{\# \text{ of mol}}{M} \\ 35.45 & & \\ \hline 58.44 & & \frac{0.089 \text{ mol}}{0.25 \text{ M}} = \boxed{0.36 \text{ L}} \end{aligned}$$

Part A Describing Molar Concentration

John O. Edwards



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