

# ATOMIC STRUCTURE

Name \_\_\_\_\_

An atom is made up of protons and neutrons (both found in the nucleus) and electrons (in the surrounding electron cloud). The atomic number is equal to the number of protons. The mass number is equal to the number of protons plus neutrons. In a neutral atom, the number of protons equals the number of electrons. The charge on an ion indicates an imbalance between protons and electrons. Too many electrons produces a negative charge, too few, a positive charge.

This structure can be written as part of a chemical symbol.

**Example:**

$\begin{array}{c} \text{mass} \\ \text{number} \\ \downarrow \\ 15 \\ \uparrow \\ 7 \end{array} \text{N}^{+3}$

*atomic number* (points to 7)  
*charge* (points to +3)

Always write the magnitude of the charge followed by the sign: N3+  
 7 protons  
 8 neutrons (15 - 7)  
 4 electrons

Can you find the other mistakes on this page?

Complete the following chart.

Element/ Ion	Atomic Number	Atomic Mass	Mass Number	Protons	Neutrons	Electrons
H						
H <sup>+</sup>						
$^{12}_6\text{C}$						
$^7_3\text{Li}^+$						
$^{35}_{17}\text{Cl}^-$						
$^{39}_{19}\text{K}$						
$^{24}_{12}\text{Mg}^{2+}$						
As <sup>3-</sup>						
Ag						
Ag <sup>+1</sup>						
S <sup>-2</sup>						
U						

# Parts Of The Atom Instructional Fair Inc

**Judd E. Hollander**



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