

Periodic Trends: Ionization Energy Worksheet

- What is the ionization energy of an element?
- Predict the following trends you expect to see in ionization energy:
 - As you move across a period? Why?
 - As you move down a group? Why?
- What factors influence the ease of removing an electron from an element?
- Choose an orbital where an electron would experience the highest Z_{eff} , effective nuclear charge (least shielded), and the highest first ionization energy. Circle the correct answer.

Na (3s)	Mg (3s)	Al (3p)	P (3p)	S (3p)
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- Match the following electron configuration with the appropriate ionization energy.

a) $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^1$	i) 1356 kJ/mol
b) $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2$	ii) 595 kJ/mol
c) $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6$	iii) 409 kJ/mol
- Which one of the following has the largest first ionization energy? Circle the correct answer.

Na	K	Li	Cs
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- The first ionization energies values (kJ/mol) for sequential elements in period 3 are shown below. Assign elements to these values based on the periodic trends.

740	578	786	1012	999
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- Choose the element with the greatest first ionization energy. Circle the correct answer.

a) Carbon or aluminum	b) Calcium or strontium
c) Helium or lithium	d) Chlorine or argon

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IAFC



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