Name

Class

Date

Chapter 7 Chemical Reactions

Section 7.1 Describing Reactions

This section discusses the use of chemical equations and how to balance them. It also demonstrates the use of calculations in chemistry.

Reading Strategy (page 192)

Monitoring Your Understanding Preview the Key Concepts, topic headings, vocabulary, and figures in this section. List two things you expect to learn. After reading, state what you learned about each item you listed. For more information on this Reading Strategy, see the Reading and Study Skills in the Skills and Reference Handbook at the end of your textbook.

What I Expect to Learn	What I Learned
Answers may vary. Possible answers: How to balance chemical equations	Answers may very. Possible answers: An unbelanced equation can be halanced by changing the coefficients.
How to convert from mass to moles	The mass of a substance can be converted to moles by using the motar mass as a conversion factor.

Chemical Equations (pages 192-193)

- Is the following sentence true or false? The new substances formed as a result of a chemical reaction are called products.
- Circle the letter of each sentence that is a correct interpretation of the chemical equation C + O₃ → CO₃.
 - Carbon and oxygen react and form carbon monoxide.
 - (b) Carbon and oxygen react and form carbon dioxide.
 - c. Carbon dioxide yields carbon and oxygen.
 - (d) The reaction of carbon and oxygen yields carbon dioxide.
- Is the following sentence true or false? The law of conservation of mass states that mass is neither created nor destroyed in a chemical reaction.
- 4. Circle the letter of the correct answer. According to the equation C + O₂ → CO₂, how many carbon atoms react with 14 molecules of oxygen to form 14 molecules of carbon dioxide?
 - a. 1 (c)14

- b. 7 d. 28
- In the reaction represented by the equation C + O₂ → CO₂, the mass of carbon dioxide produced equals the total mass of carbon and oxygen that reacted

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Marg Brown, Lois Brown

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photometry and atomic fluorescence spectroscopy along with their applications Chapter 7 concentrates on mass spectrometry with a detailed explanation on various sources of ionization and mass analyzers Chapter 8 pertains to various chromatographic separation procedures including paper chromatography thin layer chromatography column chromatography ion exchange chromatography gel filtration chromatography affinity chromatography high performance liquid chromatography and gas liquid chromatography Each type of chromatographic separation technique includes their basic principle instrumentation and applications Lastly Chapter 9 covers the importance and application of radioisotopes types of particles and their properties radioactive decay and disintegration rate interactions of radiations with matter radioactivity detection techniques and their instrumentation etc Each chapter of the book contains a few model questions to help the learners self assess their grasp of the subject as well as practice the frequently asked questions in various competitive examinations Necessary references have been incorporated to motivate readers for further exploration

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Control and Testing and Artifacts combines existing content with updates into these important parts of radiography

Molecular Biology of the Cell 6E - The Problems Book John Wilson, Tim Hunt, 2014-11-21 The Problems Book helps students appreciate the ways in which experiments and simple calculations can lead to an understanding of how cells work by introducing the experimental foundation of cell and molecular biology Each chapter reviews key terms tests for understanding basic concepts and poses research based problems The Problems Book has be

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