

pH AND pOH

Chapter 14 Section 14.2

The pH of a solution indicates how acidic or basic that solution is.

pH ranges of 0-7 = acidic

7 = neutral

7-14 = basic

Since $pH + pOH = 14$ for all aqueous solutions, the pOH of a solution can be calculated using either value.

$pH = 4 \Rightarrow pOH =$

So if $pH = 10$ then $pOH =$

$pOH = 4 \Rightarrow pH =$

So if $pOH = 10$ then $pH =$

Together $pH + pOH = 14$.

Complete the following chart.

	pH	pH	pOH	pOH	Acids or Bases
1.	normal	7	normal	7	Neutral
2.	10^{-12}	12	10^{-2}	2	acidic
3.	10^{-10}	10	normal	4	acidic
4.	normal	7	10^{-7}	7	neutral
5.	10^{-7}	7	10^{-7}	7	neutral
6.	10^{-6}	6	10^{-8}	8	basic
7.	10^{-5}	5	normal	9	basic
8.	normal	7	10^{-7}	7	neutral
9.	10^{-3}	3	10^{-11}	11	acidic
10.	10^{-2}	2	10^{-12}	12	acidic

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