

$$\text{pH} - \text{pOH} = [\text{H}^+] - [\text{OH}^-]$$

1. Calculate the values of both pH and pOH of the following solutions:

	pH	pOH
a. 0.020 M HCl		
b. 0.0050 M NaOH		
c. A blood sample $7.2 \times 10^{-8} \text{ M}$ of H^+		
d. 0.00035 M KOH		

2. Find the values of $[\text{H}^+]$, pOH, $[\text{OH}^-]$, that correspond to each of the following pH values:

	$[\text{H}^+]$	$[\text{OH}^-]$	pOH
a. pH of lemon juice = 2.90			
b. pH of sauerkraut = 3.85			
c. pH of milk of magnesia, a laxative = 10.81			
d. pH of most orange juices = 4.11			
e. pH of dilute household ammonia in windex = 11.61			

3. Determine which of the solutions in #2 are acidic?

4. A certain brand of rootbeer has a hydrogen concentration equal to $1.9 \times 10^{-8} \text{ M}$. What is the pH and pOH of this rootbeer?

5. Dr. Pepper has a $[\text{H}^+] = 1.4 \times 10^{-5} \text{ M}$. What is its pH?

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