

Precipitate Reactions: Chemical “Drop-Outs”

Many chemical reactions take place in solution. In this experiment, you will see some dissolved chemicals react when solutions are combined and drop out of solution or precipitate. In a chemical reaction, substances that react to form new substances are called reactants; new substances that are produced are called products. Substances that “fall out” of solution as solids, denoted $_{(s)}$, are called precipitates (ppt).

Goals:

- To learn about double replacement reactions.
- To learn solubility rules.
- To learn small-scale lab techniques.
- To write balanced chemical equations to describe reactions performed in lab.

Materials Needed:

- 6 small scale pipettes (empty, barium chloride, magnesium nitrate, copper (II) sulfate, potassium hydroxide, and sodium carbonate)
- reaction surface (Make sure the surface is clean, but DO NOT USE SOAP.)

Procedure:

1. Obtain materials.
2. Put 3 drops of sodium carbonate on the reaction surface in A. Add 3 drops of barium chloride solution. **Be careful to not cross contaminate the tip of the pipets.**
3. Record your observations in the data chart. Pay special attention to the color of each precipitate that forms, as well as its consistency (grainy, clumpy, etc.) . Use a Q-tip if needed to assist with your observations. Write “No Rxn” when no precipitate forms.
4. Make all of the remaining combinations listed in the data chart on the reaction surface in boxes B-J. Use 3 drops of each solution as you did in step 2. Record your observations in the data chart as soon as you make a combination.
5. Return all pipets **bulb down to the proper container** and clean the reaction surface with a paper towel.

Post-lab:

6. Complete the word equations for the reaction under part 2 in the equation chart.
7. Write the chemical equation for part 1 in the equation chart for each combination.
For your equations from part 1 in the equation chart:
8. If a precipitate formed, use the solubility chart to determine which of the products formed is the precipitate (use the solubility table). Circle the precipitate in the equation and label with $_{(s)}$. Label the other product with $_{(aq)}$.
9. If no precipitate is formed, label each product with $_{(aq)}$. Then, write “No Rxn” under the number in the pair column.
10. Answer post-lab questions.

Precipitation Chemistry Lab Answers

Khalil H. Mancy, Frederick G. Pohland



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Laboratory Guide ... General Chemistry G. W. Benton, 1895 **Chemical Laboratory Manual** Lotta Jean Bogert, 1927

A Laboratory Outline of General Chemistry Alexander Smith, 1907 **Laboratory Exercises in General Chemistry**

William Ripley Nichols, Lewis Mark Norton, 1884 *Fundamentals of Chemistry Laboratory Studies* Frank

Brescia, 2012-12-02 *Fundamentals of Chemistry Laboratory Studies* focuses on the techniques involved in chemical laboratory operations. Divided into 13 parts, the manual gives information on weights and measures, the different states of matter, atomic and molecular weights, and electron charge. Giving support to these discussions are experiments that show the changes in weight and electron charge of metals, gases, and other materials when exposed to different conditions. The text also looks at experiments on the gravimetric and volumetric stoichiometry of chlorides, sulfates, acids, antimony, and oxalates. The manual also highlights studies conducted on potassium nitrate and chlorate, oxygen, hydrogen, and polymers. The guidebook ends with discussions on molecular geometry, kinetics, and chemical equilibrium. Experiments and illustrations of chemical reactions are presented. Taking into consideration the value of data presented, the manual is a great find for readers wanting to introduce an organized system in conducting laboratory experiments. A Laboratory Course in Biomaterials Wujing Xian, 2009-06-18

The field of biomedical engineering has vastly expanded in the past two decades as reflected in the increased number of bioengineering and biomaterials programs at universities. The growth of this area has outpaced the development of laboratory courses that allow students hands-on experience since the barriers involved in creating multidisciplinary

Chemical Treatment William C. Anderson, 2013-06-29 This monograph on chemical treatment is one of a series of eight on innovative site and waste remediation technologies that are the culmination of a multi-organization effort involving more than 100 experts over a two-year period. It provides the experienced practicing professional guidance on the application of innovative processes considered ready for full-scale application. Other monographs in this series address bioremediation, soil washing, soil flushing, solvent chemical extraction, stabilization, solidification, thermal desorption, thermal destruction, and vacuum vapor extraction. **Chemical Treatment** The term chemical treatment as used in this monograph refers to the use of reagents to destroy or chemically modify target contaminants by means other than pyrolysis or combustion. The monograph addresses processes that chemically treat contaminated soils, groundwaters, surface waters, and to a limited extent, concentrated contaminants. Chemical treatment is a means of converting hazardous constituents into less environmentally objectionable forms in order to meet treatment objectives. This monograph addresses substitution, oxidation, and chemical precipitation processes. It addresses processes within these classes that are sufficiently advanced for full-scale application. There are a number of emerging technologies within these classes that are in the research or an early development stage, not yet ready for full-scale application that appear to be very promising technologically. Six such technologies are briefly addressed in Appendix A. Proceedings South African Institute of Mining and Metallurgy, 1899 **Laboratory Manual of**

General Chemistry Hermann Irving Schlesinger, Adeline De Sale Link, 1926 **Inquiry into the Singapore Science Classroom** Aik-Ling Tan, Chew-Leng Poon, Shirley S.L. Lim, 2014-05-19 This book offers an insight into the research and practices of science teaching and learning in the Singapore classroom with particular attention paid to how they map on to science as inquiry It provides a spectrum of Singapore's science educational practices through all levels of its education system detailing both successes and shortcomings The book features a collection of research and discourse by science educators in Singapore organized around four themes that are essential components of approaching science as inquiry teachers ideas and their practices opportunities and constraints from a systemic level students competencies and readiness to learn through inquiry and the need for greater awareness of the role of informal learning avenues in science education In addition the discourse within each theme is enriched by commentary from a leading international academic which helps to consolidate ideas as well as position the issues within a wider theoretical and international context Overall the papers set out important contexts for readers to understand the current state of science education in Singapore They also highlight strengths and gaps in practices of science as inquiry as well as provide suggestions about how the system can be improved These research findings are therefore helpful as they provide honest and evidence based feedback as well as tangible and doable ideas that policy makers teachers students and school administrators can adopt adapt and enhance

Chemical News and Journal of Industrial Science, 1775 **Toxicology Research Projects Directory**, 1979 **Core Science Lab Manual with Practical Skills for Class IX** V. K. Sally, Chhaya Srivastava, Goyal Brothers Prakashan, 2019-01-01 Goyal Brothers Prakashan **A Laboratory Outline of Intermediate Chemistry** Alexander Smith, 1919 **Quantitative Analysis Laboratory Manual** Dr. Bupinderpal Singh, Dr. Bikramjit Singh, Mr. Akash Bans, 2025-06-29 **Practical Skills in Science Class 09** R.P. Manchanda, Practical Book **The Chemical News and Journal of Physical Science**, 1907

A Laboratory Outline of Smith's College Chemistry Alexander Smith, 1923 **The Chemical News**, 1871

Chemical news and Journal of physical science, 1881

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