

CHAPTER 6 REVIEW*Chemical Bonding***SECTION 6-1****SHORT ANSWER** Answer the following questions in the space provided.

1. a A chemical bond between atoms results from the attraction between electrons and _____.
(a) protons (c) isotopes
(b) neutrons (d) Lewis structures
2. b A covalent bond consists of _____.
(a) a shared electron (c) two different ions
(b) a shared electron pair (d) an octet of electrons
3. a If two covalently bonded atoms are identical, the bond is identified as _____.
(a) nonpolar covalent (c) nonionic
(b) polar covalent (d) dipolar
4. b A covalent bond in which there is an unequal attraction for the shared electrons is _____.
(a) nonpolar (c) ionic
(b) polar (d) dipolar
5. c Atoms with a strong attraction for electrons they share with another atom exhibit _____.
(a) zero electronegativity (c) high electronegativity
(b) low electronegativity (d) Lewis electronegativity
6. c Bonds that possess between 5% and 50% ionic character are considered to be _____.
(a) ionic (c) polar covalent
(b) pure covalent (d) nonpolar covalent
7. a The greater the electronegativity difference between two atoms bonded together, the greater the bond's percentage of _____.
(a) ionic character (c) metallic character
(b) covalent character (d) electron sharing
8. The electrons involved in the formation of a chemical bond are called _____
valence electrons
9. A chemical bond that results from the electrostatic attraction between positive and negative ions is called a(n) _____
ionic bond

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A. Kayode Coker



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Table of Contents Prentice Hall Chemistry Chapter 11 Worksheets Answers

1. Understanding the eBook Prentice Hall Chemistry Chapter 11 Worksheets Answers
 - The Rise of Digital Reading Prentice Hall Chemistry Chapter 11 Worksheets Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Prentice Hall Chemistry Chapter 11 Worksheets Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Prentice Hall Chemistry Chapter 11 Worksheets Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Prentice Hall Chemistry Chapter 11 Worksheets Answers
 - Personalized Recommendations
 - Prentice Hall Chemistry Chapter 11 Worksheets Answers User Reviews and Ratings
 - Prentice Hall Chemistry Chapter 11 Worksheets Answers and Bestseller Lists

5. Accessing Prentice Hall Chemistry Chapter 11 Worksheets Answers Free and Paid eBooks
 - Prentice Hall Chemistry Chapter 11 Worksheets Answers Public Domain eBooks
 - Prentice Hall Chemistry Chapter 11 Worksheets Answers eBook Subscription Services
 - Prentice Hall Chemistry Chapter 11 Worksheets Answers Budget-Friendly Options
6. Navigating Prentice Hall Chemistry Chapter 11 Worksheets Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Prentice Hall Chemistry Chapter 11 Worksheets Answers Compatibility with Devices
 - Prentice Hall Chemistry Chapter 11 Worksheets Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Prentice Hall Chemistry Chapter 11 Worksheets Answers
 - Highlighting and Note-Taking Prentice Hall Chemistry Chapter 11 Worksheets Answers
 - Interactive Elements Prentice Hall Chemistry Chapter 11 Worksheets Answers
8. Staying Engaged with Prentice Hall Chemistry Chapter 11 Worksheets Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Prentice Hall Chemistry Chapter 11 Worksheets Answers
9. Balancing eBooks and Physical Books Prentice Hall Chemistry Chapter 11 Worksheets Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Prentice Hall Chemistry Chapter 11 Worksheets Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Prentice Hall Chemistry Chapter 11 Worksheets Answers
 - Setting Reading Goals Prentice Hall Chemistry Chapter 11 Worksheets Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Prentice Hall Chemistry Chapter 11 Worksheets Answers
 - Fact-Checking eBook Content of Prentice Hall Chemistry Chapter 11 Worksheets Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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