

THE NATURE OF GASES

(p1)

Review

18 pts
- 2 pt is 1st
- 1 pt is 2nd

- Describe the assumptions of the kinetic theory as it applies to gases.
- Express gas pressure in terms of kinetic theory.
- Define the relationship between Kelvin temperature and average kinetic energy.

Vocabulary

- kinetic energy
- kinetic theory
- gas pressure
- vacuum
- atmospheric pressure
- barometer

Part A Completion

Use this completion exercise to check your understanding of the concepts that are introduced in this section. Each blank can be completed with a term, short phrase, or number.

The kinetic theory describes the 1 of particles in matter and the forces of attraction between them. The theory assumes that the volume occupied by a gas is mostly 2, that the particles of gas are relatively 3, move 4 of each other, and are in constant 5 motion. The 6 between particles are perfectly elastic so that the total 7 remains constant. Gas pressure results from the simultaneous collisions of billions of particles with an object. Barometers are used to measure 8 pressure. Standard conditions are defined as a temperature of 9 and a pressure of 10.

- motion
- empty space
- far apart
- independently
- random / rapid
- collisions
- kinetic energy
- atmospheric
- 0°C / 273 K
- 1 atm / 101.3 kPa

Part B True-False

Classify each of these statements as always true, AT, sometimes true, ST, or never true, NT.

- ST 11. Atmospheric pressure is 760 mm Hg.
AT 12. The SI unit of pressure is the pascal.

↳ a multiple of pascal

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M Mark



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