

9.1 Assessment

3. The positive charge of a group A metal is equal to its group number. The charge of a group A nonmetal is the group number minus 8.
4. from the number of electrons lost
5. -ite or -ate
6. Group 1A metals, 1+; Group 3A (aluminum), 3+; Group 5A nonmetals, 3-
7. A monoatomic anion is a single atom with a negative charge; a polyatomic anion is two or more bound atoms with a negative charge.
8. a.) K^+ , cation, potassium ion
b.) O^{2-} anion, oxide ion
c.) Sn^{2+} cation, tin (II)
d.) Br^- anion, bromide ion
e.) Be^{2+} cation beryllium ion
f.) Co^{3+} cation, cobalt(III) ion
9. a.) NH_4^+
b.) Sn^{2+}
c.) CrO_4^{2-}
d.) NO_3^-

9.2 Assessment

14. Write the name of the cation followed by the name of the anion.
15. Write the symbol of the cation followed by the symbol of the anion, then use subscripts to balance the charges.
16. Write the symbol for the metal ion followed by the formula of the polyatomic ion and balance the charges. Name the cation first followed by the anion.
17. a.) $BeCl_2$
b.) Cs_2S
c.) NaI
d.) SrO
18. a.) $Cr(NO_3)_3$
b.) $NaClO_4$
c.) $Mg(HCO_3)_2$
d.) $Ca(C_2H_3O_2)_2$
19. a.) incorrect; charges are not balanced $MgSO_4$
b.)
c.) incorrect; charges are not balanced $BeCl_2$

9.3 Assessment

20. Prefixes indicate the number of atoms of each element in a molecule of the compound.

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Table of Contents Prentice Hall Chemistry Chapter 9 Assessment Answers

1. Understanding the eBook Prentice Hall Chemistry Chapter 9 Assessment Answers
 - The Rise of Digital Reading Prentice Hall Chemistry Chapter 9 Assessment Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Prentice Hall Chemistry Chapter 9 Assessment Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Prentice Hall Chemistry Chapter 9 Assessment Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Prentice Hall Chemistry Chapter 9 Assessment Answers

- Personalized Recommendations
 - Prentice Hall Chemistry Chapter 9 Assessment Answers User Reviews and Ratings
 - Prentice Hall Chemistry Chapter 9 Assessment Answers and Bestseller Lists
5. Accessing Prentice Hall Chemistry Chapter 9 Assessment Answers Free and Paid eBooks
 - Prentice Hall Chemistry Chapter 9 Assessment Answers Public Domain eBooks
 - Prentice Hall Chemistry Chapter 9 Assessment Answers eBook Subscription Services
 - Prentice Hall Chemistry Chapter 9 Assessment Answers Budget-Friendly Options
 6. Navigating Prentice Hall Chemistry Chapter 9 Assessment Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Prentice Hall Chemistry Chapter 9 Assessment Answers Compatibility with Devices
 - Prentice Hall Chemistry Chapter 9 Assessment Answers Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Prentice Hall Chemistry Chapter 9 Assessment Answers
 - Highlighting and Note-Taking Prentice Hall Chemistry Chapter 9 Assessment Answers
 - Interactive Elements Prentice Hall Chemistry Chapter 9 Assessment Answers
 8. Staying Engaged with Prentice Hall Chemistry Chapter 9 Assessment Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Prentice Hall Chemistry Chapter 9 Assessment Answers
 9. Balancing eBooks and Physical Books Prentice Hall Chemistry Chapter 9 Assessment Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Prentice Hall Chemistry Chapter 9 Assessment Answers
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Prentice Hall Chemistry Chapter 9 Assessment Answers
 - Setting Reading Goals Prentice Hall Chemistry Chapter 9 Assessment Answers
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Prentice Hall Chemistry Chapter 9 Assessment Answers

- Fact-Checking eBook Content of Prentice Hall Chemistry Chapter 9 Assessment Answers
- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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