

- B)  $1s^2 2s^2 2p^2$   
 C)  $1s^2 2s^2 2p^2 3s^2 3p^2 4s^2 3d^2$   
 D) more than one correct response  
 E) no correct response

8. Which of the following should require the smallest energy?

- A)  $1^{\text{st}}$  ionization energy of Al  
 B)  $2^{\text{nd}}$  ionization energy of Al  
 C)  $1^{\text{st}}$  ionization energy of Si  
 D)  $2^{\text{nd}}$  ionization energy of Si  
 E) Not enough information gives or none of the above

- (a)  $1s^2 2s^2 2p^6 3s^2$       (b)  $1s^2 2s^2 2p^6 3s^2$       (c)  $1s^2 2s^2 2p^6 3s^2 3p^2$   
 (d)  $1s^2 2s^2 2p^6 3s^2 3p^4$       (e)  $1s^2 2s^2 2p^6 3s^2 3p^5$

9. Of the electron configurations above, the atom with the most negative electron affinity is \_\_\_\_\_.

- A) (d)      B) (c)      C) (a)      D) (e)      E) (b)

10. Electrons in \_\_\_\_\_ bonds remain localized between two atoms. Electrons in \_\_\_\_\_ bonds can become delocalized between more than two atoms.

- A) pi, sigma      B) sigma, pi      C) pi, pi  
 D) sigma, sigma      E) ionic, sigma

11. A molecular orbital can accommodate a maximum of \_\_\_\_\_ electrons.

- A) one      B) two      C) four      D) six      E) twelve

12. There are \_\_\_\_\_ unhybridized p atomic orbitals in an sp-hybridized carbon atom.

- A) 0      B) 1      C) 2      D) 3      E) 4

13. According to valence bond theory, which orbitals overlap in the formation of the bond in HBr?

- A)  $1s$  on H and  $4p$  on Br  
 B)  $1s$  on H and  $4s$  on Br  
 C)  $1s$  on H and  $3p$  on Br  
 D)  $2s$  on H and  $4p$  on Br  
 E)  $2s$  on H and  $3p$  on Br

14. What is the formal charge on the sulfur atom of structure A?



- A) 2      B) 1      C) 0      D) -1      E) -2

# Prentice Hall Chemistry Covalent Bonding Test Answers

**Yi-Tong Ma**



**Prentice Hall Chemistry Covalent Bonding Test Answers:**

This book delves into Prentice Hall Chemistry Covalent Bonding Test Answers. Prentice Hall Chemistry Covalent Bonding Test Answers is an essential topic that needs to be grasped by everyone, from students and scholars to the general public. The book will furnish comprehensive and in-depth insights into Prentice Hall Chemistry Covalent Bonding Test Answers, encompassing both the fundamentals and more intricate discussions.

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