

**A. Choose the correct answer**

- \_\_\_\_\_ 1. Equilibrium is characterized by
  - a. opposing processes occurring at equal rates.
  - b. an open system.
  - c. net change in the amount of substance in one phase.
  - d. changes in physical states.
- \_\_\_\_\_ 2. The process of changing from a solid to a gas is known as
  - a. condensation.
  - b. sublimation.
  - c. deposition.
  - d. vaporization.
- \_\_\_\_\_ 3. Which of the following liquids is most volatile?
  - a. water
  - b. ether
  - c. molten salt
  - d. skim milk
- \_\_\_\_\_ 4. A phase is defined as any part of a system that has
  - a. the same composition throughout.
  - b. the same properties throughout.
  - c. the same properties and composition throughout.
  - d. the same temperature throughout.
- \_\_\_\_\_ 5. Dry ice sublimates into carbon dioxide gas. If the proper conditions are maintained and the system is closed, the dry ice and the carbon dioxide gas will eventually
  - a. become the same phase.
  - b. reach equilibrium.
  - c. stop changing phase.
  - d. Both (a) and (b)
- \_\_\_\_\_ 6. If the vapor pressure of water is 4.25 kPa at 30.0°C, the vapor pressure of water at 10.0°C is most likely to be
  - a. 0 kPa.
  - b. 1.23 kPa.
  - c. 4.25 kPa.
  - d. 12.34 kPa.
- \_\_\_\_\_ 7. The temperature at which the vapor pressure of a substance equals atmospheric pressure is its
  - a. boiling point.
  - b. critical point.
  - c. melting point.
  - d. triple point.
- \_\_\_\_\_ 8. The amount of energy needed to melt one mole of a substance is its molar
  - a. enthalpy of fusion.
  - b. enthalpy of vaporization.
  - c. entropy of fusion.
  - d. entropy of vaporization.

# Prentice Hall Chemistry Standardized Test Questions

## Intermolecular Forces

**Lingsheng Yao**



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