Part B True-False

_					
	8.	AT	IO. ST	12. 0	WIT:
	9.	NT	II. NT		

Part C Matching

13. a	15. b	17. d
14. 6	1.6	

Part D Questions and Problems

20.
$$\Delta H = 5.53 \frac{\text{mol NH}_4\text{NO}_2\text{UU}}{25.7 \text{ kJ}} \times \frac{25.7 \text{ kJ}}{1 \frac{\text{mol NH}_5\text{NO}_2\text{UU}}{1.000 \text{ kJ}_4\text{NO}_2\text{UU}}} = 142 \text{ kJ}$$

Section 17.4

Part A Completion

- sum
 enthalpy
- 2. enthalpy
- 3. indirectly
- 4. changed (reversed)
- 5. standard heat of formation
- 6. change
- 7. cense
- 8. 4//
- 9. pero
- 10. subtracting

Part B True-False

11. NT 13. AT 15. NT 12. NT 14. NT

Part C Matching

16. b 18. a 20. c 17. c 19. d

Part D Questions and Problems

21.
$$CuO(s) \rightarrow Cu(s) + \frac{1}{2}O_2 \Delta H = +155 \text{ kJ}$$

 $H_2(g) + \frac{1}{2}O_2(g) \rightarrow +H_2O(g) \Delta H = -242 \text{ kJ}$
 $CuO(s) + H_2(g) \rightarrow Cu(s) + H_2O(g) \Delta H_{min} = -87 \text{ kJ}$

22.
$$\Delta H_i^0$$
 (products)
= $3 \mod \Theta_i^0$ × $\frac{-393.5 \text{ kJ}}{1 \mod \Theta_i^0}$;
= -1181 kJ
 ΔH_i^0 (reactants)
= $3 \mod \Theta_i^0$ × $\frac{-110.5 \text{ kJ}}{1 \mod \Theta_i^0}$ + (-822.1 kJ)
= -1154 kJ
 $\Delta H^0 = \Delta H_i^0$ (products) $-\Delta H_i^0$ (reactants)
= $-1181 \text{ kJ} - (-1154 \text{ kJ})$
= -27 kJ

Practice Problems

Section 17.1

1.
$$200.0 \text{ Cal} \times \frac{1000 \text{ cal}}{1 \text{ Cal}} \times \frac{4.184 \text{ J}}{1 \text{ cal}} \times \frac{1 \text{ kJ}}{10^3 \text{ J}}$$

 $= 836.8 \text{ kJ}$
2. $C = \frac{525.0 \text{ cal}}{25.0 \text{ g} \times 15.0 ^{\circ}\text{C}} = 1.40 \text{ cal/(g} \times ^{\circ}\text{C)}$
3. $\Delta T = \frac{1255.0 \text{ J}}{100.0 \text{ g} \times 2.1 \text{ J/(g} \times ^{\circ}\text{C)}} = 6.0 ^{\circ}\text{C}$
4. $q = 100.0 \text{ g} \times 120.0 ^{\circ}\text{C} \times 0.90 \frac{\text{J}}{\text{g} \times ^{\circ}\text{C}}$
 $= 1.1 \times 10^3 \text{ J}$

Section 17.2

1.
$$\Delta H = 150.0 \text{ g} \times 4.18 \frac{\text{J}}{\text{g} \times \text{°C}} \times 10^{\circ}\text{C}$$

= $6.3 \times 10^3 \text{ J} = 6.3 \text{ kJ}$

2.
$$\Delta H = 15.0 \text{ g CasOH5}_2(3) \times \frac{1 \text{ mol CasOH5}_2(3)}{74.1 \text{ g CasOH5}_2(3)} \times \frac{-65.2 \text{ kJ}}{1 \text{ mol CasOH5}_2(3)} = -13.2 \text{ kJ}$$

3.
$$\Delta H = 52.4 \text{ g.CH2(S)} \times \frac{1 \text{ mol GH2(S)}}{16.0 \text{ g.CH2(S)}} \times \frac{-890.2 \text{ kJ}}{1 \text{ mol GH2(S)}} = 2.93 \times 10^3 \text{ kJ}$$

4.
$$4NH_3 + 5O_2(g) \rightarrow 4NO(g) + 6H_2O(g)$$

$$\Delta H = \frac{-226 \text{ kJ}}{1 \text{ mol NH}_2(g)} \times 4 \text{ mol NH}_2(g)$$

$$= -904 \text{ kJ}$$

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J Spring

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The Journal of Education ,1893

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years the 17 of June has much deeper significance This was the day in 1953 when we finally fled from our life of oppression which had lasted 20 years Two successive dictatorships one of Hitler and the other of Stalin caused the most horrific slaughter of civilians and soldiers eclipsing all prior history During these first years of my life I was plainly lucky to survive After this day I had a much better chance to experience the freedom needed to lead a life of creativity satisfaction and ultimately prosperity all directed largely by our own decisions th The 17 of June 1953 was a Wednesday I stayed in the apartment of my parents in my hometown of Brandenburg in the German Democratic Republic GDR the former Russian occupied zone of Germany The summer vacation of the Humboldt University in East Berlin some 40 mi further east had just started But I was alone with my father Vati Subject Guide to Books in Print ,1971 The Chemical Age ,1923

The United States Catalog Marion E. Potter,1912

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