

Answers for Predicting Products of Chemical Reactions

For all combustion reactions of hydrocarbons or hydrogen-carbon-oxygen molecules, the products will always be CO_2 and H_2O (assuming a complete reaction).



Note that Mg can only form Mg^{2+} and I can only form I^- , so the product will must have a 1:2 cation:anion ratio.



Note that the product is not H_2Cl_2 . It is important to recognize that CuCl_2 is made of three ions, Cu^{2+} and two Cl^- .



In this question, you must recognize that perchlorate, ClO_4^- , and hydroxide, OH^- , are polyatomic ions and will not break apart. Also, this is an acid-base reaction, so the products should be salt and water.



When reactions have heat as a reactant, it is very likely that they will involve decompositions. Carbonate compounds usually decompose to CO_2 and a metal oxide.



Note that one reactant is in its elemental form; if a single replacement reaction is going to occur, the species NOT in its elemental form in the reactants (H^+ in this case), will end up in its elemental form in the products (H_2) and the species that IS in its elemental form (Zn) will end up ionized (Zn^{2+}). Note that zinc can *only* form a Zn^{2+} ion, so it will have two chloride ions. Note also that hydrogen in its elemental form is H_2 , not H.



Again notice that one species is in its elemental form (Na). The magnesium in MgCl_2 is an ion (Mg^{2+}), but is transformed into its elemental state (Mg), while the Na is converted into an ion (Na^+ , sodium *only* forms a +1 ion).



Recognize that carbonate is a polyatomic ion (CO_3^{2-}) and that the cations are already stable ions (Ca^{2+} and K^+).



Note that both materials are elemental species, so the only result could be a synthesis reaction. In the product, K^+ and Cl^- are formed. Note that we would NOT write the product as K_2Cl . Ionic compounds are written in their simplest forms.



Note that phosphate (PO_4^{3-}) is a polyatomic ion and will not break apart. Since barium is a +2 ion, the barium phosphate will have a 3:2 ratio of $\text{Ba}:\text{PO}_4$ in order to balance the charge.



Notice that sulfate (SO_4^{2-}) is a polyatomic ion and that potassium exists as a +1 ion (K^+). Also, this is an acid-base reaction, so the products should be salt and water.



Another decomposition that will generate CO_2 and a metal oxide (note that Al forms a +3 ion, Al^{3+} , and monatomic oxygen will have a -2 charge, O^{2-}).

Predicting Reaction Products Answers

Therald Moeller



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Tofan,2010-07-28 This workbook is a comprehensive collection of solved exercises and problems typical to AP introductory and general chemistry courses as well as blank worksheets containing further practice problems and questions It contains a total of 197 learning objectives grouped in 28 lessons and covering the vast majority of the types of problems that a student will encounter in a typical one year chemistry course It also contains a fully solved 50 question practice test which gives students a good idea of what they might expect on an actual final exam covering the entire material *Shriver and Atkins' Inorganic Chemistry* Peter Atkins,2010 Inorganic Chemistry fifth edition represents an integral part of a student s chemistry education Basic chemical principles are set out clearly in Foundations and are fully developed throughout the text culminating in the cutting edge research topics of the Frontiers which illustrate the dynamic nature of inorganic chemistry

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