

1B-2 Practice Problems: Mass spectrometry, empirical and molecular formulae

1. Calculate percentage composition of a hydrocarbon molecule if 0.00100 mol of it weighs 0.0384 g and it yields 0.00352 g of CO_2 on combustion of 0.00100 mol.
2. Calculate the relative molecular mass of a hydrocarbon containing 85.8% carbon by mass and 14.2% hydrogen and which on combustion of 0.00100 mol yields 0.00352 g of CO_2 and 0.00144 g of H_2O .
3. Calculate percentage composition of a hydrocarbon molecule if 0.00100 mol of it weighs 0.0384 g and it yields 0.00352 g of CO_2 on combustion of 0.00100 mol.
4. Calculate the relative molecular mass of a hydrocarbon containing 85.8% carbon by mass and 14.2% hydrogen and which on combustion of 0.00100 mol yields 0.00352 g of CO_2 and 0.00144 g of H_2O .
5. Calculate percentage composition of a hydrocarbon molecule if 0.00100 mol of it weighs 0.0384 g and it yields 0.00352 g of CO_2 and 0.00144 g of H_2O on combustion of 0.00100 mol.
6. Calculate percentage composition of a hydrocarbon molecule if 0.00100 mol of it weighs 0.0384 g and it yields 0.00352 g of CO_2 and 0.00144 g of H_2O on combustion of 0.00100 mol.

Empirical and Molecular Formulae

1. Calculate the relative molecular mass of a hydrocarbon containing 85.8% carbon by mass and 14.2% hydrogen and which on combustion of 0.00100 mol yields 0.00352 g of CO_2 and 0.00144 g of H_2O .
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