

## 11-2 Practice Problems

1. Determine the mass of lithium hydroxide produced when 0.38 g of lithium nitride reacts with water according to the following equation:



$$\frac{0.38\text{g Li}_3\text{N}}{1} \times \frac{1\text{mol Li}_3\text{N}}{35\text{g Li}_3\text{N}} \times \frac{3\text{mol LiOH}}{1\text{mol Li}_3\text{N}} \times \frac{24\text{g LiOH}}{1\text{mol LiOH}} = \boxed{0.78\text{g LiOH}}$$

2. What mass of sodium chloride is produced when chlorine reacts with 0.29 g of sodium iodide?



$$\frac{0.29\text{g NaI}}{1} \times \frac{1\text{mol}}{150} \times \frac{2\text{mol NaCl}}{2\text{mol NaI}} \times \frac{58\text{g}}{1\text{mol}} = \boxed{0.112\text{g NaCl}}$$

3. Determine the mass of carbon dioxide produced when 0.85 g of butane reacts with oxygen according to the following equation:



$$\frac{0.85\text{g C}_4\text{H}_{10}}{1} \times \frac{1\text{mol}}{58\text{g}} \times \frac{8\text{mol CO}_2}{2\text{mol C}_4\text{H}_{10}} \times \frac{44\text{g}}{1\text{mol}} = \boxed{2.58\text{g CO}_2}$$

4. Determine the mass of antimony produced when 0.46 g of antimony(III) oxide reacts with carbon according to the following equation:



$$\frac{0.46\text{g Sb}_2\text{O}_3}{1} \times \frac{1\text{mol}}{291.52\text{g}} \times \frac{2\text{mol Sb}}{1\text{mol Sb}_2\text{O}_3} \times \frac{121.76\text{g}}{1\text{mol}} = \boxed{0.38\text{g Sb}}$$

5. What mass of hydrogen peroxide ( $\text{H}_2\text{O}_2$ ) must decompose to produce 0.77 g of water?



$$\frac{0.77\text{g H}_2\text{O}}{1} \times \frac{1\text{mol}}{18\text{g}} \times \frac{2\text{H}_2\text{O}_2}{2\text{H}_2\text{O}} \times \frac{34\text{g}}{1\text{mol}} = \boxed{1.45\text{g H}_2\text{O}_2}$$

6. What mass of carbon monoxide must react with oxygen to produce 0.69 g of carbon dioxide?



$$\frac{0.69\text{g CO}_2}{1} \times \frac{1\text{mol}}{44\text{g}} \times \frac{2\text{mol CO}}{2\text{mol CO}_2} \times \frac{28\text{g}}{1\text{mol}} = \boxed{0.44\text{g CO}}$$

7. Determine the mass of sodium nitrate produced when 0.73 g of nickel(II) nitrate reacts with sodium hydroxide according to the following equation:



$$\frac{0.73\text{g Ni}(\text{NO}_3)_2}{1} \times \frac{1\text{mol}}{185\text{g}} \times \frac{2\text{mol NaNO}_3}{1\text{mol Ni}(\text{NO}_3)_2} \times \frac{85\text{g}}{1\text{mol}} = \boxed{0.68\text{g NaNO}_3}$$

8. Determine the mass of calcium hydroxide produced when calcium carbide reacts with 0.64 g of water according to the following equation:



$$\frac{0.64\text{g H}_2\text{O}}{1} \times \frac{1\text{mol}}{18\text{g}} \times \frac{1\text{mol Ca}(\text{OH})_2}{2\text{mol H}_2\text{O}} \times \frac{74\text{g}}{1\text{mol}} = \boxed{1.32\text{g Ca}(\text{OH})_2}$$

9. How many grams of ozone ( $\text{O}_3$ ) must decompose to produce 0.87 g of oxygen?



$$\frac{0.87\text{g O}_2}{1} \times \frac{1\text{mol}}{32\text{g O}_2} \times \frac{2\text{mol}}{3\text{mol}} \times \frac{48\text{g O}_3}{1\text{mol}} = \boxed{0.87\text{g O}_3}$$

10. Find the mass of sugar ( $\text{C}_6\text{H}_{12}\text{O}_6$ ) required to produce 1.82 L of carbon dioxide gas at STP from the reaction described by the following equation:



11. How many liters of oxygen are necessary for the combustion of 425 g of sulfur, assuming that the reaction occurs at STP? The balanced equation is  $\text{S} + \text{O}_2 \rightarrow \text{SO}_2$ .

12. Find the mass of benzene ( $\text{C}_6\text{H}_6$ ) required to produce 2.66 L of carbon dioxide gas at STP from the reaction described by the following equation:



# Prentice Hall Chemistry Chapter 11 Assessment Answers

**Michael J. Padilla**



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*Prentice Hall Chemistry*, 2000      *Prentice Hall Health's Q and A Review of Medical Technology/clinical Laboratory Science* Anna P. Ciulla, Georganne K. Buescher, 2002 A valuable review for a wide range of laboratory professionals this book prepares candidates for certification examinations by presenting them with the latest technology and terminology as well as current test taking formats Its large number of practice questions variety of practice modes and explanations for clarification prepare learner for success on examinations Comprehensive coverage of laboratory medicine includes clinical chemistry hematology hemostasis immunology immunohematology microbiology urinalysis and body fluids molecular diagnostics laboratory calculations general laboratory principles and safety laboratory management education and computers and laboratory informatics      *Teaching and Learning in the School Chemistry Laboratory* Avi Hofstein, Muhamad Hugerat, 2021-11-05 Research into the educational effectiveness of chemistry practical work has shown that the laboratory offers a unique mode of instruction assessment and evaluation Laboratory work is an integral and important part of the learning process used to encourage the development of high order thinking and learning alongside high order learning and thinking skills such as argumentation and metacognition Authored by renowned experts in the field of chemistry education this book provides a holistic approach to cover all issues related to learning and teaching in the chemistry laboratory With sections focused on developing the skill sets of teachers as well as approaches to supporting students in the laboratory the book offers a comprehensive look at vicarious instruction methods teacher and students roles and the blend with ICT simulations and other effective approaches to practical work The book concludes with a focus on retrospective issues followed up with a look to the future of laboratory learning A product of nearly fifty years of research this book will be useful for chemistry teachers curriculum developers researchers in chemistry education and professional development providers

*Prentice Hall Science Explorer* Michael J. Padilla, 2002      *Prentice Hall Science Explorer: Chemical Interactions* Pearson/Prentice Hall, Michael J. Padilla, 2004-10      *Rodak's Hematology - E-Book* Elaine M. Keohane, Catherine N. Otto, Jeanine M. Walenga, 2019-02-22 Selected for Doody's Core Titles 2024 in Laboratory Technology Make sure you are thoroughly prepared to work in a clinical lab Rodak's Hematology Clinical Principles and Applications 6th Edition uses hundreds of full color photomicrographs to help you understand the essentials of hematology This new edition shows how to accurately identify cells simplifies hemostasis and thrombosis concepts and covers normal hematopoiesis through diseases of erythroid myeloid lymphoid and megakaryocytic origins Easy to follow and understand this book also covers key topics including working in a hematology lab complementary testing areas such as flow cytometry cytogenetics and molecular diagnostics the parts and functions of the cell and laboratory testing of blood cells and body fluid cells UPDATED nearly 700 full color illustrations and photomicrographs make it easier for you to visualize hematology concepts and show what you'll encounter in the lab with images appearing near their mentions in the text to minimize flipping pages back and forth

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**Reactor Safety Study (Rasmussen Report)** United States. Congress. House. Committee on Interior and Insular Affairs. Subcommittee on Energy and the Environment, 1976

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