

CHEMISTRY: LIMITING REAGENT Practice Problems

Please show your work on a separate piece of paper!

Assume standard temperature and pressure, STP conditions for all reactions.

1. At high temperatures, sulfur combines with iron to form the brown-black iron (II) sulfide:



In one experiment, 7.62 g of Fe are allowed to react with 8.67 g of S.

- What is the limiting reagent, and what is the reactant in excess?
- Calculate the mass of FeS formed.

2. Acrylonitrile, $\text{C}_3\text{H}_3\text{N}$, is the starting material for the production of a kind of synthetic fiber (acrylics) and can be made from propylene, C_3H_6 , by reaction with nitric oxide, NO , as follows:



What mass of $\text{C}_3\text{H}_3\text{N}$ can be made when 21.6 g of C_3H_6 react with 21.6 g of nitric oxide?

3. Calculate the percent yield for the reaction: $\text{P}_{4(s)} + 6 \text{Cl}_{2(g)} \rightarrow 4 \text{PCl}_{3(l)}$
if 75.0 g of phosphorus reacts with excess chlorine gas to produce 111.0 g of phosphorus trichloride.

4. Formic acid, HCHO_2 , burns in oxygen to form carbon dioxide and water as follows:



If a 3.15-g sample of formic acid was burned in 2.0 L of oxygen, what volume of carbon dioxide would be produced?

5. Zinc metal reacts with hydrochloric acid to produce zinc chloride and hydrogen gas.
- Balance the following reaction: $\text{Zn}_{(s)} + \text{HCl}_{(aq)} \rightarrow \text{ZnCl}_{2(aq)} + \text{H}_{2(g)}$
 - A 3.50-g sample of zinc metal is allowed to react with 2.50 g of hydrochloric acid.

Complete the following table:

reactants/products	Zn (grams)	HCl (grams)	ZnCl ₂ (grams)	H ₂ (L)
Before reaction				
After reaction	1.26 g			

Prentice Hall Chemistry Practice Problems Limiting Reactant

Theodore Lawrence Brown



Prentice Hall Chemistry Practice Problems Limiting Reactant:

Prentice Hall Chemistry, 2000 Problems and Problem Solving in Chemistry Education Georgios

Tsapralis, 2021-05-17 Problem solving is central to the teaching and learning of chemistry at secondary tertiary and post tertiary levels of education opening to students and professional chemists alike a whole new world for analysing data looking for patterns and making deductions As an important higher order thinking skill problem solving also constitutes a major research field in science education Relevant education research is an ongoing process with recent developments occurring not only in the area of quantitative computational problems but also in qualitative problem solving The following situations are considered some general others with a focus on specific areas of chemistry quantitative problems qualitative reasoning metacognition and resource activation deconstructing the problem solving process an overview of the working memory hypothesis reasoning with the electron pushing formalism scaffolding organic synthesis skills spectroscopy for structural characterization in organic chemistry enzyme kinetics problem solving in the academic chemistry laboratory chemistry problem solving in context team based active learning technology for molecular representations IR spectra simulation and computational quantum chemistry tools The book concludes with methodological and epistemological issues in problem solving research and other perspectives in problem solving in chemistry With a foreword by George Bodner

Addison-Wesley Chemistry Antony C. Wilbraham, Dennis D. Staley, Michael S. Matta, Edward L. Waterman, Prentice-Hall Staff, 2001-02 To purchase or download a workbook click on the Purchase or Download button to the left To purchase a workbook enter the desired quantity and click Add to Cart To download a free workbook right click the FREE Download PDF link and save to your computer This will result in a faster download as opposed to left clicking and opening the link

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Chemistry Charles H. Corwin, 1994 The book focuses on the concepts of chemistry and the applications that maintain and generate motivation for the subject of chemistry **Student's Guide, Chemistry, the Central Science** James C. Hill, 1991

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Modelling and Numerical Simulations II Mordechai Schlesinger, 2009-08-15 The present volume is the second in a two volume set dealing with modelling and numerical simulations in electrochemistry Emphasis is placed on the aspect of nanoelectrochemical issues It seems appropriate at this juncture to mention the growing body of opinion in some circles that George Box was right when he stated three decades ago that All models are wrong but some are useful Actually when the statement itself was made it would have been more appropriate to say that All models are inaccurate but most are useful nonetheless At present however the statement as it was made is far more appropriate and closer to the facts than ever before Currently we are in the midst of the age of massively abundant data Today's philosophy seems to be that we do not need to know why one piece of information is better than another except through the statistics of incoming and outgoing links between information and this is good enough It is why both in principle and in practice one can translate between two languages without knowledge of either While none of this can be ignored and it may even be true that All models are wrong and increasingly you can succeed without them the traditional approach of scientific modelling is still the order of the day That approach may be stated as hypothesize measure model test It is in this light that the present volume should be viewed **Chemistry** Theodore Lawrence Brown, 1988 Study Guide to

Accompany Elements of Chemistry, General, Organic, and Biological -- Robert Boikess/Kenneth Breslauer/Edward Edelson Robert A. Klein, 1986 **Atoms, Molecules, and Reactions** Ronald James Gillespie, 1994 Traditional college level chemistry including principles

Perry's Chemical Engineers' Handbook, 9th Edition Don W. Green, Marylee Z. Southard, 2018-07-13 Up to Date Coverage of All Chemical Engineering Topics from the Fundamentals to the State of the Art Now in its 85th Anniversary Edition this industry standard resource has equipped generations of engineers and chemists with vital information data and insights Thoroughly revised to reflect the latest technological advances and processes Perry's Chemical Engineers Handbook Ninth Edition provides unsurpassed coverage of every aspect of chemical engineering You will get comprehensive details on chemical processes reactor modeling biological processes biochemical and membrane separation process and chemical plant safety and much more This fully updated edition covers Unit Conversion Factors and

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Essentials of Chemical Reaction Engineering H. Scott Fogler, 2017-10-26 Today's Definitive Undergraduate Level Introduction to Chemical Reaction Engineering Problem Solving For 30 years H Scott Fogler's Elements of Chemical Reaction Engineering has been the 1 selling text for courses in chemical reaction engineering worldwide Now in Essentials of Chemical Reaction Engineering Second Edition Fogler has distilled this classic into a modern introductory level guide specifically for undergraduates This is the ideal resource for today's students learners who demand instantaneous access to information and want to enjoy learning as they deepen their critical thinking and creative problem solving skills Fogler successfully integrates text visuals and computer simulations and links theory to practice through many relevant examples This updated second edition covers mole balances conversion and reactor sizing rate laws and stoichiometry isothermal reactor design rate data collection analysis multiple reactions reaction mechanisms pathways bioreactions and bioreactors catalysis catalytic reactors nonisothermal reactor designs and more Its multiple improvements include a new discussion of activation energy molecular simulation and stochastic modeling and a significantly revamped chapter on heat effects in chemical reactors To promote the transfer of key skills to real life settings Fogler presents three styles of problems Straightforward problems that reinforce the principles of chemical reaction engineering Living Example Problems LEPs that allow students to rapidly explore the issues and look for optimal solutions Open ended problems that encourage students to use inquiry based learning to practice creative problem solving skills About the Web Site umich.edu/elements/5e/index.html The companion Web site offers extensive enrichment opportunities and additional content including Complete PowerPoint slides for lecture notes for chemical reaction engineering classes Links to additional software including Polymath MATLAB Wolfram Mathematica AspenTech and COMSOL Multiphysics Interactive learning resources linked to each chapter including Learning Objectives Summary Notes Web Modules Interactive Computer Games Computer Simulations and Experiments Solved Problems FAQs and links to LearnChemE Living Example Problems that provide more than 75 interactive simulations allowing students to explore the examples and ask what if questions Professional Reference Shelf containing advanced content on reactors weighted least squares experimental planning laboratory reactors pharmacokinetics wire gauze reactors trickle bed reactors

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