

**SECTION A**

**Section - A consists of Multiple Choice type questions of 1 mark each.**

**1.** The ratio of *HCF* to *LCM* of the least composite number and the least prime number is :

- (a) 1:2
- (b) 2: 1
- (c) 1: 1
- (d) 1: 3

**Answer: (a)**

**Explanation:**

The least composite number is 4, which has prime factors 2 and 2. The least prime number is 2.

The HCF (Highest Common Factor) of 4 and 2 is 2, because that is the largest number that divides both 4 and 2.

The LCM (Least Common Multiple) of 4 and 2 is 4, because that is the smallest number that is a multiple of both 4 and 2.

Therefore, the ratio of HCF to LCM of 4 and 2 is 2:4 or 1:2.

So the answer is (a) 1:2.

**2.** The roots of the equation  $x^2 + 3x - 10 = 0$  are :

- (a) 2, - 5
- (b) - 2, 5
- (c) 2, 5
- (d) - 2, - 5

**Answer: (a)**

**Explanation:**

Given,  $x^2 + 3x - 10 = 0$

$$x^2 + 5x - 2x - 10 = 0$$

$$x(x + 5) - 2(x + 5) = 0$$

$$(x + 5)(x - 2) = 0$$

$x = 2, - 5$  are the roots of the given equation.

# Papermathematics Midyear Question Paper

**Sebastian Brünink**



## **Papermathematics Midyear Question Paper:**

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 $2 \text{ NaCl} + \text{NaCl} \rightarrow 3 \text{ Al} + 2 \text{ Al}_2\text{O}_3 + 4 \text{ Al}_2\text{O}_3 + 2 \text{ Al}_2\text{O}_3 + 4 \text{ N}_2 + 2 \text{ NH}_3 + \text{N}_2$

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 $\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O}$   
at the molecular level the reaction would look something like this notice that there are two oxygen atoms on the

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web oct 13 2023 this means that you will need to balance the carbon atoms first 5 use a coefficient to balance the single carbon atom add a coefficient to the single carbon atom on the right of the equation to balance it with the 3 carbon atoms on the left of the equation  
 $\text{C}_3\text{H}_8 + 2 \text{ H}_2\text{O} \rightarrow 3 \text{ CO}_2$

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