- How do the bond angles in CH<sub>4</sub>, NH<sub>3</sub> and H<sub>2</sub>O compare?
  - A.  $CH_4 = NH_3 = H_2O$
  - B.  $CH_4 < NH_3 < H_2O$
  - C.  $NH_3 < CH_4 < H_2O$
  - D.  $H_2O$  <  $NH_3$  <  $CH_4$

#### 2.

The electronegativities of four different elements are given below (the letters are not their chemical symbols).

Element	W	X	Y	Z
Electronegativity	0.9	1.2	3.4	4.0

Based on this information which statement is correct?

- A. W is a non-metal.
- W and X form an ionic compound.
- C. Y is a metal.
- D. Y and Z form a covalent compound.

#### 3.

Which properties of the alkali metals decrease going down group 1?

- A. First ionization energy and reactivity
- B. Melting point and atomic radius
- Reactivity and electronegativity
- D. First ionization energy and melting point

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