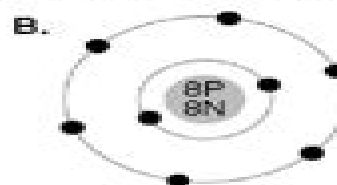
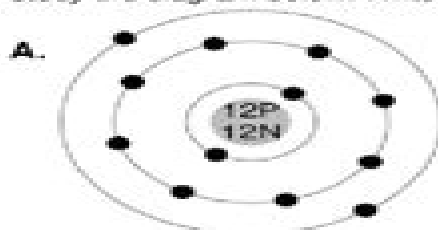


Name: _____ Period: _____

IPS Unit 9 – Chemical Bonds

Types of Bonds Worksheet

Directions: Study the diagram below. Write your answers to the questions in the spaces provided.



1. If atom A loses electrons to atom B.
 - a. how many electrons will atom A lose? _____
 - b. how many electrons will atom B gain? _____
 - c. what will be the charge of atom A? _____
 - d. what will be the charge of atom B? _____
 - e. what will be the total charge of the compound formed?

 - f. what type of bond will form?

2. Explain why an element's charge is related to the group on the periodic table to which it belongs.

Directions: Complete the table comparing ionic compounds and covalent compounds.

Characteristic	Ionic compounds	Covalent compounds
How the compound is formed		
Smallest particle		
Usual state at room temperature		

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V. ІU`Gankin, Yurii V. Gankin



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Chemical Bonding in Solids and Fluids Marcus Frederick Charles Ladd, 1994 Written by a highly regarded scientist and teacher this book examines and discusses the nature of and properties associated with interatomic and intermolecular forces in solids and fluids All discussions feature mathematical treatments accessible to chemistry students Provides stereoscopic diagrams of three dimensional structure covers computer methods where appropriate classifies and discusses solids in terms of bond type considers liquid structure and properties For students in chemistry chemical physics and biochemistry

Chemical Bonds and Bonds Energy R Sanderson, 2012-12-02 Chemical Bonds and Bonds Energy Second Edition provides information pertinent to the fundamental aspects of contributing bond energy and bond dissociation energy This book explores the values that are useful in the interpretation of significant phenomena such as product distribution and reaction mechanisms Organized into 12 chapters this edition begins with an overview of the quantitative relationship among three basic properties of an atom namely nonpolar covalent radius electronegativity and homonuclear single covalent bond energy This text then examines the quantitative means of evaluating the partial atomic charges that result from initial differences in the electromagnetivity of atoms that form a compound Other chapters consider the recognition of the reduction of bond weakening not by multiplicity and in certain types of single covalent bonds The final chapter deals with the application of the principal ideas and techniques to the oxidation of ethane This book is a valuable resource for organic and inorganic chemists

Chemical Bonds Phillip Manning, 2009 Modern life is made up of a mind boggling array of materials A simple drinking cup for example might be made of Styrofoam paper or glass depending on the drinkers needs at the moment Home storage cabinets can be made of metal wood or plastic Space shuttles are assembled from silicon steel and hundreds of other materials All of these items owe their properties to the chemical bonds between the atoms that make up the substance Chemical Bonds examines the nature of the chemical bonds answering fundamental questions about how they form how they are broken and how they help define life as we know it

Chemical Bonding Clarified Through Quantum Mechanics George C. Pimentel, Richard D. Spratley, 1969

Holt Science and Technology Holt Rinehart & Winston, Holt, Rinehart and Winston Staff, 2004-01-01

Chemical Bonds and Bonds Energy Robert Thomas Sanderson, 1976-06-28 Chemical Bonds and Bonds Energy Second Edition provides information pertinent to the fundamental aspects of contributing bond energy and bond dissociation energy This book explores the values that are useful in the interpretation of significant phenomena such as product distribution and reaction mechanisms Organized into 12 chapters this edition begins with an overview of the quantitative relationship among three basic properties of an atom namely nonpolar covalent radius electronegativity and homonuclear single covalent bond energy This text then examines the quantitative means of evaluating the partial atomic charges that result from initial differences in the electromagnetivity of atoms that form a compound Other chapters consider the recognition of the reduction of bond weakening not by multiplicity and in certain types of single covalent bonds The final

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