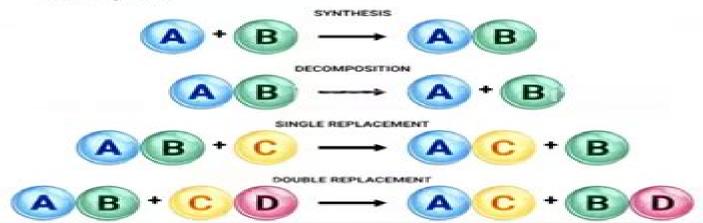
III. Types of Chemical Reactions

The 4 primary types of chemical reactions are:

- Combination (synthesis) reaction: Two or more reactants combine to make a new product. Going from simple to complex molecules.
- Decomposition reaction: A single reactant breaks down to form two or more products. Going from complex to simple molecules.
- Single Displacement reaction: A single element replaces a similar element of an adjacent reactant compound.
- Double Displacement reaction: Two compounds exchange elements, producing two new compounds.



ACTIVITY

Identify the type of reaction for the following equations.

Chemical Reaction	Type of reaction
$Cu_{(x)} + 2AgNO_{3(3q)} \rightarrow Cu(NO_3)_{2(3q)} + 2Ag_{(x)}$	
2KI _(sq) + Pb(NO ₃) _{2(sq)} → 2KNO _{3(sq)} + PbI _{2(x)}	
2H ₂ O ₂₍₀ → 2H ₂ O ₍₀₎ + O ₂₍₀₎	
C ₆ H ₁₂ O ₆ - 2C ₂ H ₅ OH + 2CO ₂	
Mg _(x) + CuSO _{4(x0)} MgSO _{4(x0)} + Cu _(x)	
2C ₈ H ₁₈ + 25O ₂ → 16CO ₂ + 18H ₂ O	

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