

Chemistry – Unit 5 Review

1. Definitions:

- acid-base equilibrium:** a chemical reaction where acid and base react to form water and a salt.
- acid rain:** rain that has a pH of less than 5.6.
- acidic solution:** a solution with a pH of less than 7.
- alkaline solution:** a solution with a pH of greater than 7.
- base:** a substance that can accept protons or donate hydroxide ions.
- base-catalyzed reaction:** a reaction where a base is used to speed up the reaction.
- buffer solution:** a solution that resists changes in pH when small amounts of acid or base are added.
- conjugate acid-base pair:** two species that differ by only one proton.
- equilibrium constant:** a value that indicates the extent of a chemical reaction.
- hydrolysis:** a reaction where a compound reacts with water.
- indicator:** a substance that changes color at a certain pH.
- neutralization:** a reaction where an acid and a base react to form water and a salt.
- pH:** a measure of the acidity or basicity of a solution.
- strong acid:** an acid that completely dissociates in water.
- strong base:** a base that completely dissociates in water.

2. Write the balanced chemical equations:

- $\text{HCl} + \text{NaOH} \rightarrow \text{NaCl} + \text{H}_2\text{O}$
- $\text{H}_2\text{SO}_4 + 2\text{NaOH} \rightarrow \text{Na}_2\text{SO}_4 + 2\text{H}_2\text{O}$
- $\text{HNO}_3 + \text{KOH} \rightarrow \text{KNO}_3 + \text{H}_2\text{O}$
- $\text{CH}_3\text{COOH} + \text{NaOH} \rightarrow \text{CH}_3\text{COONa} + \text{H}_2\text{O}$
- $\text{H}_2\text{CO}_3 + \text{NaOH} \rightarrow \text{NaHCO}_3 + \text{H}_2\text{O}$
- $\text{H}_2\text{O} + \text{CO}_2 \rightarrow \text{H}_2\text{CO}_3$
- $\text{H}_2\text{O} + \text{NH}_3 \rightarrow \text{NH}_4\text{OH}$
- $\text{H}_2\text{O} + \text{SO}_2 \rightarrow \text{H}_2\text{SO}_3$
- $\text{H}_2\text{O} + \text{NO}_2 \rightarrow \text{HNO}_3 + \text{HNO}_2$
- $\text{H}_2\text{O} + \text{P}_2\text{O}_5 \rightarrow \text{H}_3\text{PO}_4$
- $\text{H}_2\text{O} + \text{Cl}_2 \rightarrow \text{HCl} + \text{HOCl}$
- $\text{H}_2\text{O} + \text{Br}_2 \rightarrow \text{HBr} + \text{HOBr}$
- $\text{H}_2\text{O} + \text{I}_2 \rightarrow \text{HI} + \text{HOI}$
- $\text{H}_2\text{O} + \text{F}_2 \rightarrow \text{HF} + \text{HOF}$
- $\text{H}_2\text{O} + \text{O}_3 \rightarrow \text{H}_2\text{O}_2 + \text{O}_2$

3. Complete the following table with the appropriate values:

Substance	Chemical Formula	Conjugate Acid	Conjugate Base
Hydrochloric acid	HCl	H_2Cl^+	Cl^-
Sulfuric acid	H_2SO_4	H_3SO_4^+	HSO_4^-
Nitric acid	HNO_3	H_2NO_3^+	NO_3^-
Carbonic acid	H_2CO_3	H_3CO_3^+	HCO_3^-
Acetic acid	CH_3COOH	$\text{CH}_3\text{COOH}_2^+$	CH_3COO^-

4. Write the balanced chemical equations for the following reactions:

Hydrochloric acid reacts with sodium hydroxide:



- Hydrochloric acid reacts with sodium hydroxide to form sodium chloride and water.

Hydrochloric acid reacts with sodium hydroxide to form sodium chloride and water. The balanced chemical equation is:

$$\text{HCl} + \text{NaOH} \rightarrow \text{NaCl} + \text{H}_2\text{O}$$

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