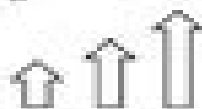


Level Up Review

Transformations

Transformations Level Up Review

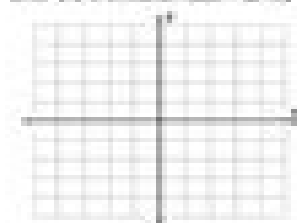
Directions: This activity has five levels of challenges. Work with 2-3 friends to solve each level. Start at Level 1. When you finish, show your work to your teacher. If it's correct, you'll move on to the next level. You have successfully completed the challenge when you finish Level 5.



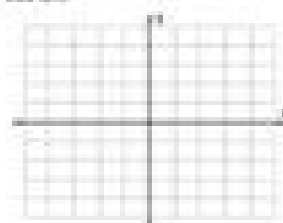
Level 1 - 2020

1. The point $A(4, -2)$ is translated by the rule $(x, y) \rightarrow (x+4, y+3)$. What are the coordinates of the image?

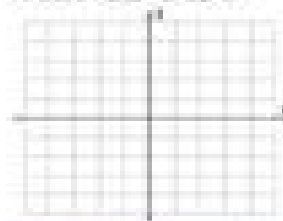
2. Find the coordinates of the image after the translation $(x, y) \rightarrow (x-1, y+2)$.



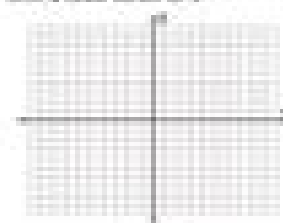
3. Graph point A with vertices $A(4, 2)$, $B(1, 5)$, and $C(3, 7)$ and the image of the image after a reflection in the y -axis.



4. Graph $ABCD$ with vertices $A(4, 2)$, $B(1, 5)$, $C(3, 7)$, and $D(2, 4)$. What is the image after the translation $(x, y) \rightarrow (x+2, y+3)$?



5. Graph quadrilateral $ABCD$ with vertices $A(4, 2)$, $B(1, 5)$, $C(3, 7)$, and $D(2, 4)$ and the image after a dilation with a scale factor of 2.



6. Describe a sequence of transformations that maps quadrilateral $ABCD$ to quadrilateral $A'B'C'D'$.



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Geometry

Tying Up Transformations Pi

Michael N. Barber, B. W. Ninham



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Table of Contents Tying Up Transformations Pi

1. Understanding the eBook Tying Up Transformations Pi
 - The Rise of Digital Reading Tying Up Transformations Pi
 - Advantages of eBooks Over Traditional Books
2. Identifying Tying Up Transformations Pi
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Tying Up Transformations Pi
 - User-Friendly Interface
4. Exploring eBook Recommendations from Tying Up Transformations Pi
 - Personalized Recommendations
 - Tying Up Transformations Pi User Reviews and Ratings
 - Tying Up Transformations Pi and Bestseller Lists
5. Accessing Tying Up Transformations Pi Free and Paid eBooks
 - Tying Up Transformations Pi Public Domain eBooks
 - Tying Up Transformations Pi eBook Subscription Services
 - Tying Up Transformations Pi Budget-Friendly Options
6. Navigating Tying Up Transformations Pi eBook Formats

- ePub, PDF, MOBI, and More
- Tying Up Transformations Pi Compatibility with Devices
- Tying Up Transformations Pi Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Tying Up Transformations Pi
 - Highlighting and Note-Taking Tying Up Transformations Pi
 - Interactive Elements Tying Up Transformations Pi
- 8. Staying Engaged with Tying Up Transformations Pi
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Tying Up Transformations Pi
- 9. Balancing eBooks and Physical Books Tying Up Transformations Pi
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Tying Up Transformations Pi
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Tying Up Transformations Pi
 - Setting Reading Goals Tying Up Transformations Pi
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Tying Up Transformations Pi
 - Fact-Checking eBook Content of Tying Up Transformations Pi
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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 $2 \text{ CO}_2(\text{g}) + \text{H}_2\text{O}(\text{g}) \rightarrow \text{C}_2\text{H}_2(\text{g}) + \frac{5}{2} \text{O}_2(\text{g}) \quad \Delta H = 94.5 \text{ kJ}$
 $\text{H}_2\text{O}(\text{g}) + \frac{1}{2} \text{O}_2(\text{g}) \rightarrow \text{H}_2\text{O}(\text{l}) \quad \Delta H = 71.2 \text{ kJ}$
 $\text{C}_2\text{H}_6(\text{g}) + \frac{7}{2} \text{O}_2(\text{g}) \rightarrow \text{CO}_2(\text{g}) + 3 \text{H}_2\text{O}(\text{g}) \quad \Delta H = 283 \text{ kJ}$
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 $\text{C}_2\text{H}_4(\text{g}) + 3 \text{O}_2(\text{g}) \rightarrow 2 \text{CO}_2(\text{g}) + 2 \text{H}_2\text{O}(\text{l}) \quad \Delta H = 1411 \text{ kJ/mole}$
 $\text{C}_2\text{H}_6(\text{g}) + \frac{7}{2} \text{O}_2(\text{g}) \rightarrow 2 \text{CO}_2(\text{g}) + 3 \text{H}_2\text{O}(\text{l}) \quad \Delta H = 1560 \text{ kJ/mole}$
 $\text{H}_2(\text{g}) + \frac{1}{2} \text{O}_2(\text{g}) \rightarrow \text{H}_2\text{O}(\text{l}) \quad \Delta H = 285 \text{ kJ/mole}$

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