

Bonding Types Lab #1: Ionic vs. Covalent Solids

Teacher's Copy

HSSC Expectations

C4.3a	Solids Solids can be classified as metallic, ionic, covalent, or network covalent. These different types of solids have different properties that depend on the particles and forces found in the solid.
C4.3c	Compare the relative strengths of forces between molecules based on the melting point and boiling point of the substances.
C4.3d	Compare the strength of the forces of attraction between molecules of different elements. (For example, at room temperature, chlorine is a gas and iodine is a solid.)
C4.3e	Predict whether the forces of attraction in a solid are primarily metallic, covalent, network covalent, or ionic based upon the elements' location on the periodic table.
C4.3f	Identify the elements necessary for hydrogen bonding (N, O, F).
C4.3g	Given the structural formula of a compound, indicate all the intermolecular forces present (dispersion, dipolar, hydrogen bonding).
C4.3h	Explain properties of various solids such as malleability, conductivity, and melting point in terms of the solid's structure and bonding.
C4.3i	Explain why ionic solids have higher melting points than covalent solids. (For example, NaF has a melting point of 995°C while water has a melting point of 0°C.)

PURPOSE:

To introduce the properties of covalent and ionic compounds and to determine the differences between the two.

PRE CAUTIONS:

The melting point part of this experiment should be performed in a fume hood. Some of the heated solids will give off a terrible-smelling vapor or smoke.

MATERIALS/EQUIPMENT:

NaCl	conductivity probe(s)
Sugar (sucrose)	glass plates
KCl	Laboratory marker
KI	fume hood
camphor	
benzoic acid	
laboratory burner or hot plate	
aluminum foil	
scoopula	

Types Of Bonding In Solids Lab Answers

J Dewey



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Table of Contents Types Of Bonding In Solids Lab Answers

1. Understanding the eBook Types Of Bonding In Solids Lab Answers
 - The Rise of Digital Reading Types Of Bonding In Solids Lab Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Types Of Bonding In Solids Lab Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Types Of Bonding In Solids Lab Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Types Of Bonding In Solids Lab Answers
 - Personalized Recommendations
 - Types Of Bonding In Solids Lab Answers User Reviews and Ratings
 - Types Of Bonding In Solids Lab Answers and Bestseller Lists
5. Accessing Types Of Bonding In Solids Lab Answers Free and Paid eBooks
 - Types Of Bonding In Solids Lab Answers Public Domain eBooks
 - Types Of Bonding In Solids Lab Answers eBook Subscription Services
 - Types Of Bonding In Solids Lab Answers Budget-Friendly Options

6. Navigating Types Of Bonding In Solids Lab Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Types Of Bonding In Solids Lab Answers Compatibility with Devices
 - Types Of Bonding In Solids Lab Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Types Of Bonding In Solids Lab Answers
 - Highlighting and Note-Taking Types Of Bonding In Solids Lab Answers
 - Interactive Elements Types Of Bonding In Solids Lab Answers
8. Staying Engaged with Types Of Bonding In Solids Lab Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Types Of Bonding In Solids Lab Answers
9. Balancing eBooks and Physical Books Types Of Bonding In Solids Lab Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Types Of Bonding In Solids Lab Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Types Of Bonding In Solids Lab Answers
 - Setting Reading Goals Types Of Bonding In Solids Lab Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Types Of Bonding In Solids Lab Answers
 - Fact-Checking eBook Content of Types Of Bonding In Solids Lab Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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