The Main Types of Chemical Reactions

A chemical reaction may be regarded as the process that occurs when matter undergoes change in composition. To be able to understand chemical reactions, one needs to recognize the key types of reactions. Reactions may be classified in several different ways, namely:

Synthesis (Combination) Reactions; Decomposition Reactions; Substitution (single replacement) Reactions; Precipitation (Double replacement) Reactions; Neutralisation (acid-base) Reactions; Redox (Reduction and Oxidation) Reactions.

Type of Reaction	Explanation	General Equation (General form of change occurrences)	Example
Synthesis/ Combination	Reaction in which two or more simpler substances (elements) combine chemically to give a compound.	A+B AB	2H ₂ + O ₂ 2H ₂ O
Decomposition	Reaction in which a compound is broken down into simpler substances (elements).	AB A + B	2H ₂ O 2H ₂ + O ₂
Substitution (Single replacement)	Reaction in which an atom or group of atoms is replaced by another atom or group.	A+BC AC+B OR A+BC BA+C	Zn + 2HCl ZnCl ₂ + H ₂ OR Cl ₂ + 2NaBr 2NaCl + Br ₂
Precipitation (Double replacement)	Reaction in which a solid compound is formed when solutions of two soluble compounds are mixed.	AB + CD AD + CB	AgNO ₃ + NaCl AgCl + NaNO ₃
Neutralisation (Acid- Base)	Reaction in which an acid reacts with a base to give salt and water.	HA + BOH → H ₁ O + BA	HCI + NaOH → H ₂ O + NaCI
Redox (Reduction- Oxidation)	Reaction in which electron transfer occurs.	Reduction:A A + e Oxidation: B + e B	aCl ₂ + 2e → 2Cl (gain electrons) Na → Na' + e (loss of electrons) 2Na + Cl ₂ → 2 NaCl

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D. Epiotis, 2012-12-06 **Femtochemistry: Ultrafast Dynamics** using proven software Of The Chemical Bond (In 2 Volumes) - Volume 1 Ahmed H Zewail, 1994-09-12 These two volumes on Femtochemistry present a timely contribution to a field central to the understanding of the dynamics of the chemical bond This century has witnessed great strides in time and space resolutions down to the atomic scale providing chemists biologists and physicists with unprecedented opportunities for seeing microscopic structures and dynamics Femtochemistry is concerned with the time resolution of the most elementary motions of atoms during chemical change bond breaking and bond making on the femtosecond 10 15 second time scale This atomic scale of time resolution has now reached the ultimate for the chemical bond and as Lord George Porter puts it chemists are near the end of the race against time These two volumes cover the general concepts techniques and applications of femtochemistry Professor Ahmed Zewail who has made the pioneering contributions in this field has from over 250 publications selected the articles for this anthology These volumes begin with a commentary and a historical chronology of the milestones He then presents a broad perspective of the current state of knowledge in femtochemistry by researchers around the world and discusses possible new directions In the words of a colleague it is a must on the reading list for all of my students all readers will find this to be an informative and valuable overview The introductory articles in Volume I provide reviews for both the non experts as well as for experts in the field This is followed by papers on the basic concepts For applications elementary reactions are studied first and then complex reactions Volume I is complete with studies of solvation dynamics non reactive systems ultrafast electron diffraction and the control of chemical reactions Volume II continues with reaction rates the concept of elementary intramolecular vibrational energy redistribution IVR and the phenomena of rotational coherence which has become a powerful tool for the determination of molecular structure via time resolution. 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calculation of unimolecular rate under vibrational potential is also given along with pertinent degrees of freedom The remaining chapters explore the energy distribution functions appropriate to each system the averaging of k E and the relations between theoretical and experimental parameters Thermal reactions chemical activation systems and the theory of mass spectra are examined The last chapter is devoted to the transition state and its ambiguities This text will be of interest to gas kineticists mass spectrometrists and students and researchers working in the field of physical chemistry for Multimedia Systems Marten J. van Sinderen, Lambert J.M. Nieuwenhuis, 2003-06-30 This conference in Enschede The Netherlands is the sixth in a series of international conferences and workshops under the title Protocols for Multimedia Systems abbreviated as PROMS The first PROMS workshop took place in June 1994 in Berlin Germany followed by workshops in Salzburg Austria October 1995 and Madrid Spain October 1996 In 1997 PROMS formed a temporary alliance with Multimedia Networking a conference previously held in Aizu Japan in 1995 This led to the international conference on Protocols for Multimedia Systems Multimedia Networking PROMS MmNet that took place in Santiago Chile November 1997 Since then PROMS has been announced as an international conference although informal contacts and interactive sessions as in a workshop were retained as a desirable feature of PROMS After a gap of three years PROMS was organized in Cracow Poland October 2000 for the fifth time We consider it a challenge to make this sixth edition of PROMS as successful as the previous events The goal of the PROMS series of conferences and workshops is to contribute to scientific strategic and practical cooperation between research institutes and industrial companies in the area of multimedia protocols This is also the goal of PROMS 2001 The basic theme of this conference continues to be multimedia protocols both at the network and application level although the increasing interest in wireless mobility and quality of service as interrelated topics with relevance to multimedia are reflected in the current program University of Iowa Studies in Natural History ,1922

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