Chemistry - Unit 6 Reaction Equations Worksheet 1

Balance the following equations by inserting the proper coefficients. For selected reactions, draw *Before* and *After* particle diagrams to show the particles involved in the reaction. Be sure to provide a key.

- 1. ___C + ___H₂O \rightarrow ___CO + ___H₂
- 2. $\underline{\hspace{0.1cm}}$ MgO \rightarrow $\underline{\hspace{0.1cm}}$ Mg + $\underline{\hspace{0.1cm}}$ O₂
- 3. $Al + O_2 \rightarrow Al_2O_3$

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Before

#8

- #3 Before
- 5. ___Cl₂ + ___KI → ___KCl + ___I₂

 $_{\text{Zn}}$ + $_{\text{H}_2}$ SO₄ \rightarrow $_{\text{ZnSO}_4}$ + $_{\text{H}_2}$

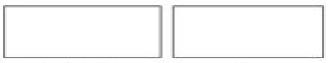
- CuCl → Cu + Cl₂
- ___Na + ___Cl₂ → ___NaCl
- 8. __Al + __HCl \rightarrow __AlCl₃ + __H₂



After

After

- 9. $_{\text{Fe}_2O_3} \rightarrow _{\text{Fe}} + _{\text{O}_2}$
- 10. ___P + ___O₂ \rightarrow ___P₂O₅
- 11. $\underline{\hspace{0.1cm}}Mg + \underline{\hspace{0.1cm}}HCl \rightarrow \underline{\hspace{0.1cm}}MgCl_{2} + \underline{\hspace{0.1cm}}H_{2}$
- 13. $__BaCl_2 + __H_2SO_4 \rightarrow __BaSO_4 + __HCl$
- 14. $__{CH_4} + __{O_2} \rightarrow __{CO_2} + __{H_2O}$



- #14 Before After
 15. a) ZnCl. + (NH.)-S → ZnS + NH.Cl
 - b) Find the molar mass of these reactants.
 - c) How many moles of ZnCl₂ would be in 25 g? How much mass would 0.55 moles of (NH₄)₂S have?

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