

Chemistry Unit 8 Worksheet Examples of Every Kind of Problem

1. **Equation:** $2\text{O}_2(g) + 2\text{H}_2\text{O}(l) \rightarrow 4\text{H}^+(aq) + 4\text{OH}^-(aq)$
Reaction: O_2 H_2O
Change: -1.2 $+1.2$ $+1.2$
Mass: 3.2 3.6
 1.2 moles and 3.6g, wanted
2. **Equation:** $\text{H}_2\text{O}_2 \rightarrow 2\text{H}^+ + \text{O}_2 + 2\text{H}_2\text{O}$
Reaction: H_2O_2 H^+ O_2 H_2O
Change: -0.0001 $+0.0002$ $+0.0001$ $+0.0002$
Mass: 0 0 0.0036 0.0036
 0.0001 moles and 0.0036g, wanted
3. **Equation:** $2\text{H}_2\text{O}_2 \rightarrow 2\text{H}_2\text{O} + \text{O}_2$
Reaction: $2\text{H}_2\text{O}_2$ $2\text{H}_2\text{O}$ O_2
Change: -0.001 $+0.001$ $+0.0005$
Mass: 0 0.036 0.009
 0.001 moles and 0.036g, wanted
4. **Equation:** $\text{H}_2\text{O}_2 + 2\text{H}^+ + 2\text{e}^- \rightarrow 2\text{H}_2\text{O}$
Reaction: H_2O_2 H^+ H_2O
Change: -0.0001 $+0.0002$ $+0.0001$
Mass: 0 0.0036 0.0036
 0.0001 moles and 0.0036g, wanted
5. **Equation:** $2\text{H}_2\text{O}_2 \rightarrow 2\text{H}_2\text{O} + \text{O}_2$
Reaction: $2\text{H}_2\text{O}_2$ $2\text{H}_2\text{O}$ O_2
Change: -0.0001 $+0.0001$ $+0.00005$
Mass: 0 0.0036 0.0009
 0.0001 moles and 0.0036g, wanted
6. **Equation:** $2\text{H}_2\text{O}_2 \rightarrow 2\text{H}_2\text{O} + \text{O}_2$
Reaction: $2\text{H}_2\text{O}_2$ $2\text{H}_2\text{O}$ O_2
Change: -0.0001 $+0.0001$ $+0.00005$
Mass: 0 0.0036 0.0009
 0.0001 moles and 0.0036g, wanted
7. **Equation:** $2\text{H}_2\text{O}_2 \rightarrow 2\text{H}_2\text{O} + \text{O}_2$
Reaction: $2\text{H}_2\text{O}_2$ $2\text{H}_2\text{O}$ O_2
Change: -0.0001 $+0.0001$ $+0.00005$
Mass: 0 0.0036 0.0009
 0.0001 moles and 0.0036g, wanted
8. **Equation:** $2\text{H}_2\text{O}_2 \rightarrow 2\text{H}_2\text{O} + \text{O}_2$
Reaction: $2\text{H}_2\text{O}_2$ $2\text{H}_2\text{O}$ O_2
Change: -0.0001 $+0.0001$ $+0.00005$
Mass: 0 0.0036 0.0009
 0.0001 moles and 0.0036g, wanted
9. **Equation:** $2\text{H}_2\text{O}_2 \rightarrow 2\text{H}_2\text{O} + \text{O}_2$
Reaction: $2\text{H}_2\text{O}_2$ $2\text{H}_2\text{O}$ O_2
Change: -0.0001 $+0.0001$ $+0.00005$
Mass: 0 0.0036 0.0009
 0.0001 moles and 0.0036g, wanted
10. **Equation:** $2\text{H}_2\text{O}_2 \rightarrow 2\text{H}_2\text{O} + \text{O}_2$
Reaction: $2\text{H}_2\text{O}_2$ $2\text{H}_2\text{O}$ O_2
Change: -0.0001 $+0.0001$ $+0.00005$
Mass: 0 0.0036 0.0009
 0.0001 moles and 0.0036g, wanted

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K Payea

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the 'K Payea' text.

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