

Unit 3: Acids and Bases

This unit activity will help you meet these essential learning goals:

- 1. **Balance Equations**—You will explore the chemical acids, bases, and neutralization reactions, including their balancing and stoichiometry and calculating pH values. You will demonstrate your knowledge by:
- 1.1. **Chemical Reactions**—You will employ stoichiometry for reactions and analysis.

Introduction

Acids and bases are common to many areas of biological and environmental systems. They make significant contributions to the products of combustion reactions. The structure of biological cells is determined by the composition and pH of the surrounding water and acid or base solutions. Substrate, catalyst, temperature, and other conditions affect the rate of chemical reactions, and studying acid-base reactions provides

- an example of how to use stoichiometry as a tool?
- an example of how to use pH as a tool?

In this activity, you will explore the relationship between pH, acids and bases, and describe the environmental impact of acids. You will also discover how different conditions can increase and decrease the rate of a reaction, and how, in turn, the rate of a reaction can affect the system's equilibrium. You will also discover how the rate of a reaction can be affected by different conditions, and how these apply to real-world examples.

Objectives and Analysis

Task 1: Acids and Bases in real world applications

Many reactions happen every day around us, such as digestion, neutralization, combustion, acid and base reactions, hydrolysis, and many others. Some reactions, such as neutralization, happen every day.

- 1. Explore the acids and bases in real world applications, and how they affect the environment and the formation of acid rain. Investigate various acids and their properties of neutralizing chemical reactions.

Open your answer book

Neutralization reaction: $\text{HCl} + \text{NaOH} \rightarrow \text{NaCl} + \text{H}_2\text{O}$
 Neutralization reaction: $\text{H}_2\text{SO}_4 + \text{NaOH} \rightarrow \text{Na}_2\text{SO}_4 + \text{H}_2\text{O}$

Plato Chemistry Unit Activity Answers

Liying Dong



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