

| AQA Chemistry (8462) from 2016 Topics C4.1 Atomic structure and the periodic table | | | | |
|---|---|---|---|---|
| Topic | Student Checklist | R | A | G |
| 4.1.1 A simple model of the atom, symbols, relative atomic mass, electronic charge and isotopes | State that everything is made of atoms and recall what they are | | | |
| | Describe what elements and compounds are | | | |
| | State that elements and compounds are represented by symbols; and use chemical symbols and formulae to represent elements and compounds | | | |
| | Write word equations and balanced symbol equations for chemical reactions, including using appropriate state symbols | | | |
| | HT ONLY: Write balanced half equations and ionic equations | | | |
| | Describe what a mixture is | | | |
| | Name and describe the physical processes used to separate mixtures and suggest suitable separation techniques | | | |
| | Describe how the atomic model has changed over time due to new experimental evidence, inc discovery of the atom and scattering experiments (inc the work of James Chadwick) | | | |
| | Describe the difference between the plum pudding model of the atom and the nuclear model of the atom | | | |
| | State the relative charge of protons, neutrons and electrons and describe the overall charge of an atom | | | |
| | State the relative masses of protons, neutrons and electrons and describe the distribution of mass in an atom | | | |
| | Calculate the number of protons, neutrons and electrons in an atom when given its atomic number and mass number | | | |
| | Describe isotopes as atoms of the same element with different numbers of neutrons | | | |
| | Define the term relative atomic mass and why it takes into account the abundance of isotopes of the element | | | |
| | Calculate the relative atomic mass of an element given the percentage abundance of its isotopes | | | |
| 4.1.2 The periodic table | Describe how electrons fill energy levels in atoms, and represent the electron structure of elements using diagrams and numbers | | | |
| | Recall how the elements in the periodic table are arranged | | | |
| | Describe how elements with similar properties are placed in the periodic table | | | |
| | Explain why elements in the same group have similar properties and how to use the periodic table to predict the reactivity of elements | | | |
| | Describe the early attempts to classify elements | | | |
| | Explain the creation and attributes of Mendeleev's periodic table | | | |
| | Identify metals and non-metals on the periodic table, compare and contrast their properties | | | |
| | Explain how the atomic structure of metals and non-metals relates to their position in the periodic table | | | |
| | Describe noble gases (group 0) and explain their lack of reactivity | | | |
| | Describe the properties of noble gases, including boiling points, predict trends down the group and describe how their properties depend on the outer shell of electrons | | | |
| | Describe the reactivity and properties of group 1 alkali metals with reference to their electron arrangement and predict their reactions | | | |
| | Describe the properties of group 7 halogens and how their properties relate to their electron arrangement, including trends in molecular mass, melting and boiling points and reactivity | | | |
| | Describe the reactions of group 7 halogens with metals and non-metals | | | |
| | Chem ONLY: Describe the properties of transition metals and compare them with group 1 elements, including melting points and densities, strength and hardness, and reactivity (for Cr, Mn Fe, Co, Ni & Cu) | | | |

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