

Answers: Oxidation Number

1. N in N_2O_3 N^{3+}

2. S in H_2SO_4 S^{6+}

3. C C^0

4. C in CO C^{2+}

5. Na in NaCl Na^{+}

6. H in H_2O H^{+}

7. Ba in BaCl_2 Ba^{2+}

8. N in NO_2^- N^{3+}

9. S in Al_2S_3 S^{2-}

10. S in HSO_4^- S^{6+}

11. Cl in $\text{Fe}(\text{ClO}_2)_3$ Cl^{3+}

12. Fe in $\text{Fe}(\text{ClO}_2)_3$ Fe^{3+}

13. N in NO_3^- N^{5+}

14. Cu²⁺ Cu^{2+}

15. Zn²⁺ Zn^{2+}

16. C in CH_4 C^{4+}

17. Mn in MnO_2 Mn^{4+}

18. S in SO_3^{2-} S^{4+}

19. Mg²⁺ Mg^{2+}

20. Cl⁻ Cl^-

21. O₂ O_2^0

22. P₄ P_4^0

23. Na in Na_2S Na^{+}

24. S in H_2S S^{2-}

25. Ca²⁺ Ca^{2+}

26. C in CN^- C^{4+}

27. H in OH^- H^{+}

28. Mn in KMnO_4 Mn^{7+}

29. I in $\text{Mg}(\text{IO}_3)_2$ I^{5+}

30. C in $\text{C}_2\text{O}_4^{2-}$ C^{2+}

Oxidation Numbers Section Review Answers Answers

John T. Moore, Richard H. Langley



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