

Name : _____

Precipitation Reactions and Solubility Worksheet

For each pair of compounds, predict the reaction that will occur and write equations to describe the reaction. Use the solubility rules to help you predict the state (solid or aqueous) of the products. Hint: In each reaction, a precipitate will be formed, and be sure to balance the equations if necessary.

Example: $\text{NaCl (aq)} + \text{AgNO}_3 \text{ (aq)}$

Molecular Equation:



Complete Ionic Equation: Split each aqueous component into its ions.



Net Ionic Equation: Eliminate any spectator ions which do not undergo chemical change.



Word Equation:

Aqueous sodium chloride is added to aqueous silver nitrate to form solid silver chloride plus aqueous sodium nitrate.

2) $\text{AgNO}_3 \text{ (aq)} + \text{Na}_2\text{CrO}_4 \text{ (aq)}$

Molecular Equation:

Complete Ionic Equation:

Net Ionic Equation:

Word Equation:

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