

## Revision - Principles of chemistry 1e - Chemical formulae and calculations

Where do you find the mass number on the periodic table? The big number

How do you calculate the mass of a NaOH: Add the top numbers together -  $23+16+1 = 40$

a)  $\text{F}_2$  \_\_\_\_\_

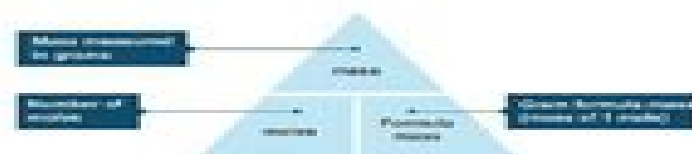
c)  $\text{H}_2\text{SO}_4$  \_\_\_\_\_

e)  $\text{Mg}(\text{OH})_2$  \_\_\_\_\_

b)  $\text{Fe}$  \_\_\_\_\_

d)  $\text{Al}_2\text{O}_3$  \_\_\_\_\_

f)  $\text{Al}(\text{NO}_3)_3$  \_\_\_\_\_



Calculate the number of moles in

a) 100 g of  $\text{Cu}$  \_\_\_\_\_

b) 1 kg of  $\text{Al}_2\text{O}_3$  \_\_\_\_\_

c) 20 mg of  $\text{Au}$  \_\_\_\_\_

d) 0.35g  $\text{HNO}_3$  \_\_\_\_\_

Calculate the mass of

a) 5 moles of  $\text{N}_2$  \_\_\_\_\_

b) 0.02 moles of  $\text{Al}_2\text{O}_3$  \_\_\_\_\_

c) 2 moles of  $\text{NaCl}$  \_\_\_\_\_

e) 0.05 moles of  $\text{NaCl}$  \_\_\_\_\_

d) 0.2 moles  $\text{C}_2\text{H}_6$  \_\_\_\_\_

### Percentage yield

A chemistry teacher burned 50g of copper in oxygen, hoping to make 80g of reaction only produced 67g of this product. What was the percentage yield?

copper oxide. However, the

$$\% \text{ Yield} = \frac{\text{Actual Yield}}{\text{Predicted Yield}} \times 100$$

# Preliminary Chemistry Revision Questions

**Fudeko Reekie**



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