

14

THE BEHAVIOR OF GASES

Chapter Test II

A. Matching

Match a word or phrase in Column I with the correct description in Column II. Write the letter of the correct answer next to the item.

Column I

1. The constant relating unit temperature changes to changes in volume for a fixed amount of gas at constant pressure (the ratio of the change in volume to the change in temperature).
2. The volume of a fixed amount of a gas at constant pressure in the standard state (0°C and 1 atm).
3. The ratio of the mass of a gas to its volume, independent of the actual amount of gas used.
4. An experimental gas constant determined from the ideal gas law at 0°C and 1 atm.
5. A constant of nature that has values of 8.314 J/mol·K or 0.08206 L·atm/mol·K.
6. The a parameter of gas constant temperatures for determining the gas constant from the ideal gas law.
7. An experimental gas constant determined from the ideal gas law at 0°C and 1 atm.
8. $\frac{P_1 V_1}{T_1} = \frac{P_2 V_2}{T_2}$
9. The amount of heat enough to raise 1 g of water 1°C.
10. A unit of energy (calorie).

Column II

- a. standard gas law
- b. ideal gas constant R
- c. volume
- d. compressibility
- e. Boyle's law
- f. molar volume
- g. standard molar volume
- h. density
- i. Joule-calorie
- j. heat capacity of water

B. Multiple Choice

Choose the correct answer and write its letter on the line.

1. Expanding one volume of a monatomic gas from 1 atm and 0°C to 2 atm and 0°C requires a change in:
a. the temperature from 0°C to 273°C
b. the volume to two times
c. the pressure to two times
d. the volume to two times

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Scott C. Dulebohn



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