

Study Guide

Key Concepts



14.1 Properties of Gases

- Gases are easily compressed because of the space between particles in a gas.
- The amount of gas (n), volume (V), and temperature (T) are factors that affect gas pressure (P).



14.2 The Gas Laws

- As the pressure of a gas increases, the volume decreases, if the temperature is constant.
- As the temperature of an enclosed gas increases, the volume increases, if the pressure is constant.
- As the temperature of an enclosed gas increases, the pressure increases, if the volume is constant.

- The combined gas law allows you to do calculations for situations in which only the amount of gas is constant.



14.3 Ideal Gases

- To calculate the number of moles of a contained gas requires an expression that contains the variable n .
- Real gases differ most from an ideal gas at low temperatures and high pressures.



14.4 Gases: Mixtures and Movements

- In a mixture of gases, the total pressure is the sum of the partial pressures of the gases.
- Gases of lower molar mass diffuse and effuse faster than gases of higher molar mass.

Vocabulary

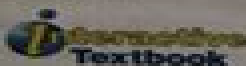
- Boyle's law (p. 418)
- Charles's law (p. 420)
- combined gas law (p. 424)
- compressibility (p. 413)
- Dalton's law of partial pressures (p. 432)
- diffusion (p. 435)
- effusion (p. 435)
- Gay-Lussac's law (p. 422)
- Graham's law of effusion (p. 435)
- ideal gas constant (p. 426)
- ideal gas law (p. 426)
- partial pressure (p. 432)

Key Equations

- Boyle's law: $P_1 \times V_1 = P_2 \times V_2$
- Charles's law: $\frac{V_1}{T_1} = \frac{V_2}{T_2}$
- Gay-Lussac's law: $\frac{P_1}{T_1} = \frac{P_2}{T_2}$
- Combined gas law: $\frac{P_1 \times V_1}{T_1} = \frac{P_2 \times V_2}{T_2}$
- Ideal gas law: $P \times V = n \times R \times T$ or $PV = nRT$
- Dalton's law: $P_{\text{total}} = P_1 + P_2 + P_3 + \dots$
- Graham's law: $\frac{\text{Rate}_A}{\text{Rate}_B} = \sqrt{\frac{\text{molar mass}_B}{\text{molar mass}_A}}$

Organizing Information

Use these terms to construct a concept map that organizes the major ideas of this chapter.



Concept Map 14 Solve the Concept Map with the help of an interactive guided tutorial.

with **ChemASAP**

Boyle's law

Charles's law

partial pressure

combined gas law

diffusion

Prentice Hall Chemistry Section 14 Lab Answers

Michael Hoelscher



Prentice Hall Chemistry Section 14 Lab Answers:

This book delves into Prentice Hall Chemistry Section 14 Lab Answers. Prentice Hall Chemistry Section 14 Lab Answers is an essential topic that must be grasped by everyone, ranging from students and scholars to the general public. This book will furnish comprehensive and in-depth insights into Prentice Hall Chemistry Section 14 Lab Answers, encompassing both the fundamentals and more intricate discussions.

1. The book is structured into several chapters, namely:

- Chapter 1: Introduction to Prentice Hall Chemistry Section 14 Lab Answers
- Chapter 2: Essential Elements of Prentice Hall Chemistry Section 14 Lab Answers
- Chapter 3: Prentice Hall Chemistry Section 14 Lab Answers in Everyday Life
- Chapter 4: Prentice Hall Chemistry Section 14 Lab Answers in Specific Contexts
- Chapter 5: Conclusion

2. In chapter 1, the author will provide an overview of Prentice Hall Chemistry Section 14 Lab Answers. This chapter will explore what Prentice Hall Chemistry Section 14 Lab Answers is, why Prentice Hall Chemistry Section 14 Lab Answers is vital, and how to effectively learn about Prentice Hall Chemistry Section 14 Lab Answers.
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6. In chapter 5, the author will draw a conclusion about Prentice Hall Chemistry Section 14 Lab Answers. The final chapter will summarize the key points that have been discussed throughout the book.
- The book is crafted in an easy-to-understand language and is complemented by engaging illustrations. This book is highly recommended for anyone seeking to gain a comprehensive understanding of Prentice Hall Chemistry Section 14 Lab Answers.

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