FLAME TESTS

INTRODUCTION

According to the Bohr theory of the atom, electrons may occupy only specific energy levels. When an atom absorbs sufficient energy, an electron can "jump" to a higher energy level. Higher energy levels tend to be less stable, however, and if a lower energy level is available, the electron will "fall" back, giving off energy in the process. The difference in energies between the two levels is emitted in the form of a photon of electromagnetic radiation. The energy of each photon is described by the equation E = hv, where h is Planck's constant and v is the frequency of the radiation. If the wavelength of the released photon is between 400 and 700 nm, the energy is emitted as visible light. The color of the light depends on the specific energy change that is taking place.

White light is a continuous spectrum in which all wavelengths of visible light are present. An excited atom, however, produces one or more specific lines in its spectrum, corresponding to the specific changes in energy levels of its electrons. Because each element has a distinct electron configuration, each has a unique line spectrum.

Flame tests are a quick method of producing the characteristic colors of metallic ions, the loosely-held electrons of a metal are easily excited in the flame of a lab burner. The emission of energy in the visible portion of the spectrum as those electrons return to lower energy levels produces a colored flame, the color is a combination of the wavelengths of each transition, and may be used to determine the identity of the ion.

In this investigation you will perform flame tests on seven metallic ions, then use your results to determine the identity of several unknowns.

PRE-LAB DISCUSSION

Read the entire laboratory investigation and the relevant pages of your textbook. then answer the questions that follow.

 Write out the electron configurations for each of the metallic ions to be tested in this investigation.

Ba²*
Cu²*
Li*

K* Sr²⁺

Ca²⁺ Na⁺

What does a flame test indicate about the energy changes taking place among the electrons in a metallic ion?

Prentice Hall Flame Test Lab Answers

Robert R. Redfield

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environmen especially for the fetus the infant the very young tal factors in the routine medical history The role of the elderly and the chronically ill those most sus enhanced susceptibility as a factor in an individual s ceptible Environmental medicine fOCUSing on the response to toxic exposure is discussed *Nuclear Science Abstracts*, 1964-04 <u>Chemical Abstracts</u>, 1924 **Good Lighting and the Illuminating Engineer**, 1911 **The Illuminating Engineer**, 1912

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