



• EXAMPLES: OF BASES:



• LIMITATIONS:

Following are the limitations of Arrhenius concept:

- 1) It applies to the Aqueous solution.
- 2) It doesnot explain why some compounds such as CO_2 , SO_2 are acids and NH_3 is base.

Q: Define and explain the Bronsted-Lowery concept of acids and bases with examples?

• BRONSTED - LOWERY CONCEPT OF ACIDS AND BASES:

• HISTORY:

In 1923, J.N Bronsted and T.M Lowery independently proposed another theory to overcome the shortcomings of Arrhenius theory. This theory is known as Bronsted-Lowery theory.

• DEFINITIONS:

"An acid is a proton donor and a base is a proton acceptor."

• EXPLANATION WITH EXAMPLES:

Consider the following reactions:



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M Lipman



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