

Study Guide

Key Concepts



14.1 Properties of Gases

- Gases are easily compressed because of the space between particles in a gas.
- The amount of gas (n), volume (V), and temperature (T) are factors that affect gas pressure (P).



14.2 The Gas Laws

- As the pressure of a gas increases, the volume decreases, if the temperature is constant.
- As the temperature of an enclosed gas increases, the volume increases, if the pressure is constant.
- As the temperature of an enclosed gas increases, the pressure increases, if the volume is constant.

- The combined gas law allows you to do calculations for situations in which only the amount of gas is constant.



14.3 Ideal Gases

- To calculate the number of moles of a contained gas requires an expression that contains the variable n .
- Real gases differ most from an ideal gas at low temperatures and high pressures.



14.4 Gases: Mixtures and Movements

- In a mixture of gases, the total pressure is the sum of the partial pressures of the gases.
- Gases of lower molar mass diffuse and effuse faster than gases of higher molar mass.

Vocabulary

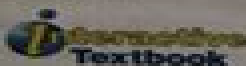
- Boyle's law (p. 418)
- Charles's law (p. 420)
- combined gas law (p. 424)
- compressibility (p. 413)
- Dalton's law of partial pressures (p. 432)
- diffusion (p. 435)
- effusion (p. 435)
- Gay-Lussac's law (p. 422)
- Graham's law of effusion (p. 435)
- ideal gas constant (p. 426)
- ideal gas law (p. 426)
- partial pressure (p. 432)

Key Equations

- Boyle's law: $P_1 \times V_1 = P_2 \times V_2$
- Charles's law: $\frac{V_1}{T_1} = \frac{V_2}{T_2}$
- Gay-Lussac's law: $\frac{P_1}{T_1} = \frac{P_2}{T_2}$
- Combined gas law: $\frac{P_1 \times V_1}{T_1} = \frac{P_2 \times V_2}{T_2}$
- Ideal gas law: $P \times V = n \times R \times T$ or $PV = nRT$
- Dalton's law: $P_{\text{total}} = P_1 + P_2 + P_3 + \dots$
- Graham's law: $\frac{\text{Rate}_A}{\text{Rate}_B} = \sqrt{\frac{\text{molar mass}_B}{\text{molar mass}_A}}$

Organizing Information

Use these terms to construct a concept map that organizes the major ideas of this chapter.



Concept Map 14 Solve the Concept Map with the help of an interactive guided tutorial.

with **ChemASAP**

Boyle's law

Charles's law

partial pressure

combined gas law

diffusion

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A Gutmann



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Ralph A. Burns, 2003 For one semester preparatory chemistry courses or general purpose introductory chemistry courses This clearly written well illustrated versatile textbook provides thorough coverage of chemistry with a balance of problem solving skills real world applications and an emphasis on critical thinking and the process of science A supporting theme throughout the text continually emphasizes that chemistry is everywhere Confronting Global Climate Change Mark Harris, 2019-06-07 This book offers a solutions based approach to climate change problems which potentially impinge on human beings within the tropics It largely comprises research articles with supplementary applications and illustrations The effects of atmospheric phenomena energy acquisition wind power CO₂ sequestration are linked with soils aquatic life reducing deforestation rainwater harvesting and clay pot farming climate plant disease and food security to show that no area of life is untouched by the phenomenon of climate change It discusses specific problem areas and provides an overview of geotechnical and sustainable solutions to lessen the impact of climate World Ocean Assessment Alan Simcock, 2017-04-17 This United Nations report examines the current state of knowledge of the world's oceans for policymakers and provides a reference for marine science courses *Clinical Biochemistry* John S. Varcoe, 2001 Clinical biochemistry is an analytical and interpretative science The analytical part involves the determination of the level of chemical components in body fluids and tissues The interpretative part examines these results and uses them in the diagnosis of disease the screening for susceptibility to specific diseases and the monitoring of the progress of treatment This book is designed to cover the major techniques and analytical instruments used in clinical biochemistry Each chapter of this book is based on a specific technique or techniques with associated instrumentation These are discussed in some detail A historical introduction is included for most of the techniques and the current uses of the techniques are presented Following that is a series of practical exercises The first exercises in most of the chapters are a general introduction to the technique leading to those with a clinical bias Where applicable the clinical practical exercises are associated with a case history and or the discussion of the relevance of the assay to diagnosis and prognosis and to the monitoring of recovery Each chapter concludes with a selection of appropriate references **The Chemical News**, 1906 *Chemical News and Journal of Industrial Science*, 1906 **Modern Enolate Chemistry** Manfred Braun, 2016-02-23 Authored by one of the world's leading synthetic chemists in the field this reference presents modern enolate chemistry with an emphasis on metal O enolates in asymmetric synthesis While great care is taken to cover novel successful concepts such classical methods as the famous Evans enolates are equally highlighted Throughout the book representative reaction procedures are presented thus helping readers to find the best solution for their own synthetic problem Of high interest to synthetic chemists in academia as well as the pharmaceuticals agrochemicals and fine chemicals industries Chemical News and Journal of Physical Science, 1906 The Illustrated London News, 1881

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