

16**SOLUTIONS****SECTION 16.1 PROPERTIES OF SOLUTIONS (pages 471–477)**

This section describes the factors that affect the solubility of a substance and determines the rate at which it occurs dissolves.

► Solution Formation (pages 471–472)

Look at Figure 16.1 on page 471 to help you answer Questions 1 and 2.

- Underline the conditions that cause sugar to dissolve faster in water.
 - as a whole cube or as granulated sugar
 - when allowed to melt or when stirred
 - at a higher temperature or at lower temperature
- Name three factors that influence the rate at which a solute dissolves in a solvent.
 - _____
 - _____
 - _____
- In the following sentence, place an **all** or **few** particles dissolve more rapidly than larger particles because their particles expose a greater surface area to the colliding solvent molecules. _____

► Solubility (pages 472–473)

- Complete the following table, showing the steps in a procedure to determine the total amount of sodium chloride that will dissolve in 100 g of water at 25°C.

Procedure	Amount Dissolved	Amount Not Dissolved
Add 35.0 g of sodium chloride to the water.	35.0 g	0.0 g
Add an additional 1.0 g of sodium chloride.		
Determine the total amount that dissolved.		

- The amount of a substance that dissolves in a given quantity of solvent at a constant temperature is called the substance's _____ at that temperature.
- If a solution contains the maximum amount of solute for a given quantity of solvent at a constant temperature, it is called a(n) _____ solution.

Prentice Hall Chemistry Section 9 Packet Answers

Robert R. Redfield



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