

## 14

## THE BEHAVIOR OF GASES

## Chapter Test II

## A. Matching

Match a word or phrase in Column I with the correct description in Column II. Write the letter of the correct answer next to the word.

## Column I

1. The constant relating unit temperature changes to changes in volume for a fixed amount of gas at constant pressure (the slope of the straight line on a P versus V graph).
2. The volume of a fixed amount of gas at standard temperature and pressure (STP).
3. The ratio of the mass of a gas to its volume, independent of the actual mass or volume used.
4. An experimental gas constant determined from the ideal gas law at STP.
5. A constant of nature that has values of about 8.314 J/mol·K.
6. The a parameter of gas constant equations for determining the gas molar mass of a gas.
7. An experimental gas constant derived from the ideal gas law at STP.
8.  $\frac{P_1 V_1}{T_1} = \frac{P_2 V_2}{T_2}$
9. The amount of heat enough to raise 1 g of water 1°C.
10. A unit of energy (Joules)

## Column II

- a. standard gas law
- b. ideal gas constant R
- c. density
- d. molar volume
- e. Boyle's law
- f. molar volume
- g. standard molar volume
- h. volume
- i. Joules/mole
- j. Joules/mole of oxygen

## B. Multiple Choice

Choose the correct answer and write its letter on the line.

1. According to Graham's law, the rate of effusion of a gas is proportional to the square root of its molar mass. Which of the following gases effuses most rapidly?  
a. the slowest of the four  
b. the one having the least mass  
c. the one having the most mass  
d. the slowest of the four

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**James E. Girard, James Girard**



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